

Like Dissolves Like Practice Answers

Using your knowledge of the "like dissolves like" principle of solubility, determine which component will dissolve best in each of the following solvents. Explain your reason for choosing each like dissolves.

1) Solvent: methanol (CH₃OH)

Components: carbon tetrachloride and oxygen dichloride

Both oxygen dichloride and methanol are polar, while carbon tetrachloride is not.

2) Solvent: carbon dioxide

Components: methanol (CH₃OH) and carbon dioxide

While both components are gases, methanol is considerably more polar than carbon dioxide (both components always have that charge separation, while carbon dioxide, no matter how polar, it's not).

3) Solvent: water

Components: carbon tetrachloride and ammonia

Both water and ammonia are polar. Carbon tetrachloride is not.

4) Solvent: hexagonal alcohol

Components: water and methanol

Both water and hexagonal alcohol are polar. Methanol is not.

Like Dissolves Like Practice Answer Key

DP Hallahan

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

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