

Moorpark College
 Chemistry 11 Mock Final Examination **ANSWERS**
 Professor Torres

- What is the pH of a solution that is $1.7 \times 10^{-2} \text{ M H}^+$?
 A. **4.77**
 B. 10.98
 C. 0.22
 D. 9.23
 E. 1.70
- Which combination below will form a buffer solution?
 A. HCl and NaCl
 B. HNO_3 and NaNO_3
 C. **$\text{HC}_2\text{H}_3\text{O}_2$ and $\text{NaC}_2\text{H}_3\text{O}_2$**
 D. NaBr and NaOH
 E. All of the above are true.
- A sample of helium gas initially at 37.0°C , 785 torr, and 2.00 L was heated to 58.0°C while the volume expanded to 3.24 L. What is the final pressure in atm?
 A. 517
 B. **0.681**
 C. 1.79
 D. 3.21
 E. None of the above
- Consider the reaction of an aqueous solution of lead(II) nitrate with an aqueous solution of potassium iodide. Which of the following represents the correctly balanced net ionic equation?
 A. $\text{Pb}^{2+}(\text{aq}) + \text{I}_2(\text{aq}) \rightarrow \text{PbI}_2(\text{s})$
 B. $2\text{NO}_3^-(\text{aq}) + 2\text{K}^+(\text{aq}) \rightarrow 2\text{KNO}_3$
 C. $\text{Pb}^{2+}(\text{aq}) + 2\text{NO}_3^-(\text{aq}) + 2\text{K}^+(\text{aq}) + 2\text{I}^-(\text{aq}) \rightarrow \text{PbI}_2(\text{s}) + 2\text{K}^+(\text{aq}) + 2\text{NO}_3^-(\text{aq})$
 D. **$\text{Pb}^{2+}(\text{aq}) + 2\text{I}^-(\text{aq}) \rightarrow \text{PbI}_2(\text{s})$**
 E. None of the above
- What is the molecular geometry of sulfur dioxide, SO_2 ?
 A. **bent**
 B. linear
 C. tetrahedral
 D. trigonal pyramidal
 E. octahedral

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