

CHEM 131 Final Exam Review Worksheet

Based on problems from the ACS Exam Review booklet, published ACS Chemistry Olympiad Exams, OpenStax Chemistry, and other sources. Note: The final exam may cover topics not included on this list or in the practice problems.

Part I	Part II	Part IV
<p><i>Introduction to Chemistry</i></p> <p><i>Scientific Method</i></p> <p><i>Measurements</i> significant figures percent yield percent error</p> <p><i>Atoms</i> isotopes atomic number mass number symbol protons, neutrons, electrons average atomic mass from isotopes</p> <p><i>Molecules</i> percent composition counting atoms/molecules</p> <p><i>Concentration</i> molarity molality percent by mass (w/w)% percent by mass-volume (w/v)% percent by volume (v/v)%</p> <p><i>Chemical Reactions</i> balancing limiting coefficients summing coefficients classification</p> <p><i>Stoichiometry</i> g → mol → percent yield molarity → mol → percent yield actual vs. percent yield titrations</p>	<p><i>Gas Laws</i> standard conditions (STP) $PV = nRT$ $R = 0.082057 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}$ $R = 8.3145 \text{ J}/\text{mol}\cdot\text{K}$</p> <p><i>Enthalpy</i> calorimetry $q = mc_p\Delta T$</p> <p><i>Hess's Law</i> standard conditions, ΔH° $\Delta H_{\text{rxn}} = \Delta H_1 + \Delta H_2 + \dots$ $\Delta H = \sum \Delta H_f^{\circ} \text{ products} - \sum \Delta H_f^{\circ} \text{ reactants}$ $\Delta H_{\text{bonds}} = \sum \Delta H_{\text{bonds broken}} - \sum \Delta H_{\text{bonds formed}}$</p>	<p><i>The Structure of the Atom</i></p> <p><i>Energy</i> $E = h\nu = hc/\lambda$ $c = \lambda\nu$ $h = 6.6260693 \times 10^{-34} \text{ J}\cdot\text{s}$ $c = 2.99792458 \times 10^8 \text{ m/s}$</p> <p><i>The Bohr Model</i> $\frac{1}{\lambda} = R_b \left(\frac{1}{n_f^2} - \frac{1}{n_i^2} \right)$ $\Delta E = k \left(\frac{1}{n_i^2} - \frac{1}{n_f^2} \right)$</p> <p><i>Quantum Numbers</i> n: principal l: angular momentum m_l: magnetic m_s: spin</p> <p><i>Electron Configuration</i> atoms vs. ions</p> <p><i>Periodic Trends</i> Z_{eff}: effective nuclear charge atomic size ionic size electronegativity electron affinity</p> <p><i>Bonding Models</i> Lewis structures valence electrons resonance structures formal charges VSEPR and geometry polarity</p> <p><i>Valence Bond Theory</i> hybrid atomic orbitals multiple bonds (π and σ) bond angles</p>
	<p>Part III</p> <p><i>Intermolecular Forces</i> dispersion dipole-dipole hydrogen bonds donors vs. acceptors induced-dipole like-dissolves-like</p> <p><i>Colligative Properties</i> vapor pressure lowering $\Delta P_{\text{solvent}} = X_{\text{solvent}} P^{\circ}_{\text{solvent}}$ $P_{\text{solvent}} = X_{\text{solvent}} P^{\circ}_{\text{solvent}}$ boiling point elevation $\Delta T_b = iK_b m$ freezing point depression $\Delta T_f = iK_f m$ osmotic pressure $\Pi = iMRT$</p>	

Montgomery College Chemistry Final Exam Review

Wenbin Ji

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular gradient shape partially visible behind it.

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Resources in Education ,1995 **Register of the University of California** University of California, Berkeley,1881

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Charles Edward Groves,1900 **Research in Education** ,1974 **Journal of the American Medical Association**

American Medical Association,1925 *Guide to American Graduate Schools* ,1990 Jocelyn Wanda Palmer,2012-10

Jocelyn Meyers met Mark Denton a gentleman navy officer from New Mexico while she s at a high school graduation party at Penny s Club with her close friend Marcie Rogers Little did she know her acquaintance with him will not only spark the beginning of a beautiful friendship but will also change both of their lives forever In so many ways Jocelyn and Mark were opposites He was shy She was gregarious He needed time to himself She found people interesting and exhilarating but did not mind being alone with her own thoughts He liked to grow things and usually had a plant in his quarters Mark took Jocelyn and her three brothers Matthew James and Robert to New Mexico to show them his world as a Navajo On the other hand Mark wanted to find out if Jocelyn could accept his family the same way he accepted hers He had fallen in love with Jocelyn and wanted her to be his wife But she never thought of their relationship as more than friendship Did she really want to spend the rest of her life with Mark What was it like to be in love with a Navajo man Was she ready to commit to him

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The Educational review). ,1891 **The Medical Directory for 1873 and General Medical Register Including the**

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- Chapter 1: Introduction to Montgomery College Chemistry Final Exam Review
- Chapter 2: Essential Elements of Montgomery College Chemistry Final Exam Review
- Chapter 3: Montgomery College Chemistry Final Exam Review in Everyday Life
- Chapter 4: Montgomery College Chemistry Final Exam Review in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Montgomery College Chemistry Final Exam Review. This chapter will explore what Montgomery College Chemistry Final Exam Review is, why Montgomery College Chemistry Final Exam Review is vital, and how to effectively learn about Montgomery College Chemistry Final Exam Review.
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6. In chapter 5, this book will draw a conclusion about Montgomery College Chemistry Final Exam Review. This chapter will summarize the key points that have been discussed throughout the book.
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