

of the following

Arsenic Fluoride
 Potassium Oxide
 Magnesium Sulfide
 Magnesium Bromide
 Magnesium Oxide
 Calcium Iodide
 Sodium Iodide
 Beryllium Chloride
 Calcium Phosphide
 Potassium Nitride
 Iron(II) Bromide
 Iron(II) Bromide
 Sodium Cyanide
 Sodium Phosphide
 Sodium Phosphide
 Magnesium Nitride
 Sodium Nitride
 Calcium Carbonate
 Calcium Hydroxide
 Aluminum Sulfate
 Tin(II) Sulfate
 Tin(IV) Sulfate
 Copper(II) Chlorate

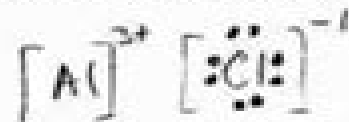
Ionic Compounds Review Sheet

Write the formulas for the following compounds:

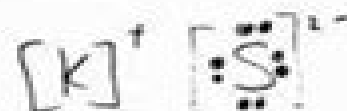
Potassium Chloride KCl
 Calcium Oxide CaO
 Magnesium Nitride Mg_3N_2
 Calcium Chloride $CaCl_2$
 Lithium Fluoride LiF
 Beryllium Phosphide Be_3P_2
 Barium Iodide BaI_2
 Strontium Chloride $SrCl_2$
 Lithium Oxide Li_2O
 Lithium Nitride Li_3N
 Magnesium Iodide MgI_2
 Aluminum Bromide $AlBr_3$
 Lead (II) Sulfide PbS
 Iron (II) Iodide FeI_2
 Lead (IV) Phosphate $Pb_4(PO_4)_3$
 Iron (III) Bromide $FeBr_3$
 Tin (IV) Fluoride SnF_4
 Tin (II) Nitride Sn_3N_2
 Copper (I) Bromide $CuBr$
 Copper (II) Sulfate $CuSO_4$
 Iron (II) Sulfite $FeSO_3$
 Ammonium Carbonate $(NH_4)_2CO_3$

Draw the Lewis Dot Diagrams for the following ionic pair:

Aluminum and chlorine



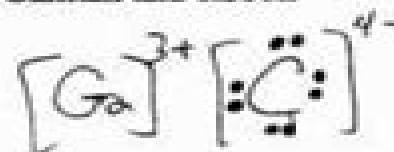
Potassium and sulfur



Zinc and phosphorus



Gallium and carbon



Monatomic Ion 1 Answers

Charles Henrickson



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