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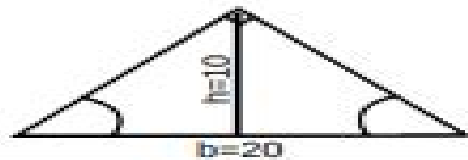
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## Geometry

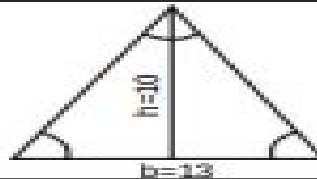
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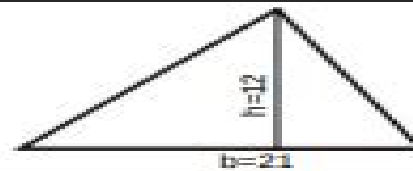
❖ Find the area of the triangle given that the area is: base X height  
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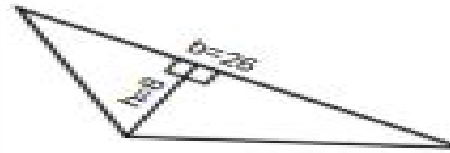
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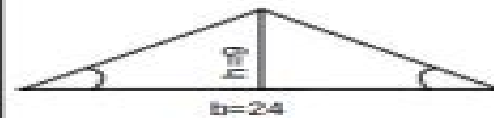
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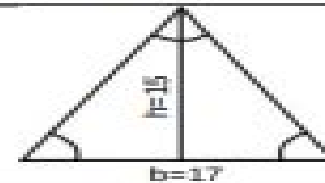
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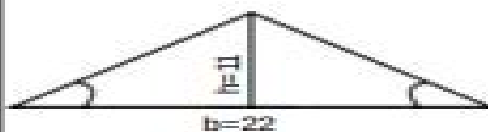
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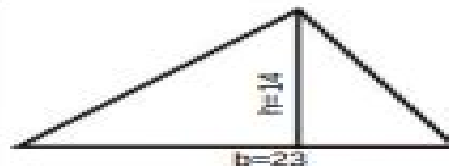
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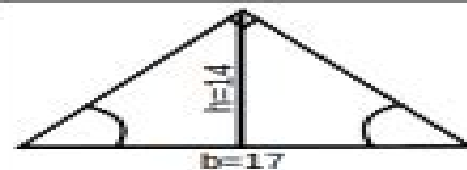
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web  $m \times c \times \Delta t$  equation 1 where  $m$  is the total mass of the solution solute plus solvent  $c$  is the specific heat of the solution and  $\Delta t$  is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of *guided inquiry designing the ideal hand warmer* - Dec 27 2022

web the heat of reaction  $\Delta H_{\text{soln}}$  is written after the products in units of  $\text{kJ mol rxn}$  for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as  $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$   $\Delta H_{\text{soln}} = -44.2 \text{ kJ mol rxn}$  write equations to similarly represent the dissolving process for

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web hand warmers are usually placed inside the gloves to keep your hands warm the heat keeping your hands warm is formed as a result of an exothermic reaction which occurs in a small packet hand warmers have 5 main ingredients in them salt carbon polypropylene vermiculite water and iron

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web oct 9 2021 sodium acetate coin or other small piece of metal plastic bags seal the coin inside a small bag place this bag inside a larger bag containing supersaturated sodium acetate get this by dissolving as much sodium acetate as possible in hot water this chemical hand warmer is reusable

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution  $\Delta H_{\text{soln}}$  or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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