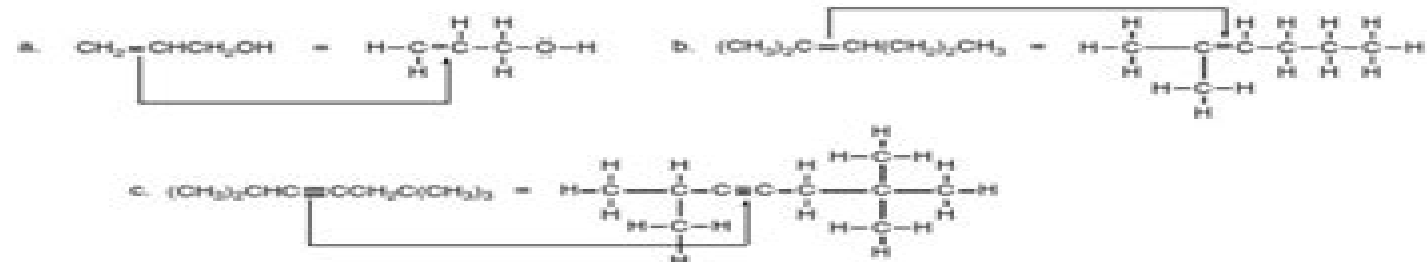


Chapter 13 Unsaturated Hydrocarbons

Solutions to In-Chapter Problems

- 13.1** To draw a complete structure for each condensed structure, first draw in the multiple bonds. Then draw in all the other C's and H's, as in Example 13.1.



- 13.2** To determine whether each molecular formula corresponds to a saturated hydrocarbon, alkene, or alkyne, recall that the formula for a saturated hydrocarbon is $\text{C}_n\text{H}_{2n+2}$, the formula for an alkene is C_nH_{2n} , and the formula for an alkyne is $\text{C}_n\text{H}_{2n-2}$.

- a. $\text{C}_5\text{H}_8 = \text{C}_n\text{H}_{2n} = \text{alkene}$ c. $\text{C}_8\text{H}_{14} = \text{C}_n\text{H}_{2n-2} = \text{alkyne}$
 b. $\text{C}_5\text{H}_{12} = \text{C}_n\text{H}_{2n+2} = \text{saturated hydrocarbon}$ d. $\text{C}_6\text{H}_{12} = \text{C}_n\text{H}_{2n} = \text{alkene}$

- 13.3** Use the general formulas [saturated hydrocarbon ($\text{C}_n\text{H}_{2n+2}$), alkene (C_nH_{2n}), and alkyne ($\text{C}_n\text{H}_{2n-2}$)] to determine the molecular formula for each compound.

- a. alkene = C_nH_{2n} , $4 \times 2 = 8$, C_4H_8
 b. saturated hydrocarbon = $\text{C}_n\text{H}_{2n+2}$, $(6 \times 2) + 2 = 14$, C_6H_{14}
 c. alkyne = $\text{C}_n\text{H}_{2n-2}$, $(7 \times 2) - 2 = 12$, C_7H_{12}
 d. alkene = C_nH_{2n} , $5 \times 2 = 10$, C_5H_{10}

- 13.4** Give the IUPAC name for each compound using the following steps, as in Example 13.2:

- [1] Find the longest chain containing both carbon atoms of the multiple bond.
 [2] Number the chain to give the double bond the lower number.
 [3] Name and number the substituents and write the complete name.



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