

- Q5.** Translate the following statements into chemical equations and then balance them.
- Hydrogen gas combines with nitrogen to form ammonia.
 - Hydrogen sulphide gas burns in air to give water and sulphur dioxide.
 - Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate.
 - Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.

Ans. (a) $3\text{H}_2(\text{g}) + \text{N}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$
 (b) $2\text{H}_2\text{S}(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\ell) + 2\text{SO}_2(\text{g})$
 (c) $3\text{BaCl}_2(\text{aq}) + \text{Al}_2(\text{SO}_4)_3(\text{aq}) \rightarrow 2\text{AlCl}_3(\text{aq}) + 3\text{BaSO}_4(\text{s})$
 (d) $2\text{K}(\text{s}) + 2\text{H}_2\text{O}(\ell) \rightarrow 2\text{KOH}(\text{aq}) + \text{H}_2(\text{g})$

- Q6.** Balance the following chemical equations.

- $\text{HNO}_3 + \text{Ca}(\text{OH})_2 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{H}_2\text{O}$
- $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
- $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{NaNO}_3$
- $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + \text{HCl}$

Ans. (a) $2\text{HNO}_3 + \text{Ca}(\text{OH})_2 \rightarrow \text{Ca}(\text{NO}_3)_2 + 2\text{H}_2\text{O}$
 (b) $2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
 (c) $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{NaNO}_3$
 (d) $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{HCl}$

- Q7.** Write the balanced chemical equations for the following reactions.

- Calcium hydroxide + Carbon dioxide $\rightarrow$ Calcium carbonate + Water
- Zinc + Silver nitrate $\rightarrow$ Zinc nitrate + Silver
- Aluminium + Copper chloride $\rightarrow$ Aluminium chloride + Copper
- Barium chloride + Potassium sulphate $\rightarrow$ Barium sulphate + Potassium chloride

Ans. (a) $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$
 (b) $\text{Zn} + 2\text{AgNO}_3 \rightarrow \text{Zn}(\text{NO}_3)_2 + 2\text{Ag}$
 (c) $2\text{Al} + 3\text{CuCl}_2 \rightarrow 2\text{AlCl}_3 + 3\text{Cu}$
 (d) $\text{BaCl}_2 + \text{K}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{KCl}$

- Q8.** Write the balanced chemical equation for the following and identify the type of reaction in each case.

- Potassium bromide(aq) + Barium iodide(aq) $\rightarrow$ Potassium iodide(aq) + Barium bromide(s)
- Zinc carbonate(s) $\rightarrow$ Zinc oxide(s) + Carbon dioxide(g)
- Hydrogen(g) + Chlorine(g) $\rightarrow$ Hydrogen chloride(g)
- Magnesium(s) + Hydrochloric acid(aq) $\rightarrow$ Magnesium chloride(aq) + Hydrogen(g)

Ans. (a) $2\text{KBr}(\text{aq}) + \text{BaI}_2(\text{aq}) \rightarrow 2\text{KI}(\text{aq}) + \text{BaBr}_2(\text{s})$;Double displacement reaction

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