# CHAPTER 14: COVALENT BONDING: ORBITALS

# PROBLEMS WITH LEWIS STRUCTURE

- · Provides no indication of the snapes of the molecules
- · Oz would be predicted to be diamagnetic, but it is paramagnetic

Poramagnetic = has unpoined electrons.

The Lewis structure of O<sub>2</sub> has an O+O double band

with no unpaired electrons, i.e. diamagnetic.

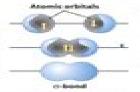
- · Many compounds cannot be represented by a single structure
- · Resonance structures differ only in position of electrons not the nuclei.
- Result in notions of hypervalent compounds and "expansion of octet"
   Can't explain molecules such as B2H6 (electron deficient molecules)

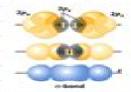
# VALENCE—BOND THEORY "LOCALIZED ELECTRON MODEL

· Covalent bond is created by the overlap of orbitals on two atoms

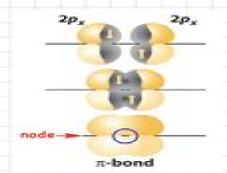
# TYPFS OF OVERLAP: Of and IT Bonds

- · Bonding in H2
  - · sausage snaped orbital, bond called o-orbital
  - · Called of orbital because it has no mode and is cylindrically symmetrical
  - · Electron defisity is concentrated between two nuclei, but is still highest at the nuclei themselves.





- · Bondina in Na
  - · or orbital is obtained by overlap of two pr-orbitals
- · Perpendicular p orbitals overlap in a side-by-side marmer.
- . The resulting orbital has two lobes of electron density and is called a 11-bond.
- · called TT-orbital because it has a node and, when viewed down the internuclear axis, it looks like a p-orbital
- · IT-bond has no electron density between the nuclei along the bond axis.



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**M Lipman** 

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