

## Electron Configuration Practice Worksheet **KEY**

*In the space below, write the unabbreviated electron configurations of the following elements:*

- 1) sodium       $1s^2 2s^2 2p^6 3s^1$
- 2) iron       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium       $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^4$

*In the space below, write the abbreviated electron configurations of the following elements:*

- 6) cobalt       $[Ar] 4s^2 3d^7$
- 7) silver       $[Kr] 5s^2 4d^9$
- 8) tellurium       $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium       $[Rn] 7s^2$
- 10) lawrencium       $[Rn] 7s^2 5f^{14} 6d^1$

Determine what elements are denoted by the following electron configurations:

- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  **Sulfur**
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  **Rubidium**
- 13)  $[Kr] 5s^2 4d^{10} 5p^3$  **Antimony**
- 14)  $[Xe] 6s^2 4f^{14} 5d^8$  **Osmium**
- 15)  $[Rn] 7s^2 5f^{11}$  **Einsteinium**

Determine which of the following electron configurations are not valid:

- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$       **should be  $3d^{10}$**
- 17)  $1s^2 2s^2 2p^6 3s^2 3d^5$        **$3s^2$  is impossible & "p" instead of "d"**
- 18)  $[Ra] 7s^2 5f^8$       **Ra is not a noble gas**
- 19)  $[Kr] 5s^2 4d^{10} 5p^5$       **Valid**
- 20)  $[Xe]$       **Element can't be its own electron configuration.**

# Electron Configurations Worksheet Answer Key

**Y Pai**



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## **Table of Contents Electron Configurations Worksheet Answer Key**

1. Understanding the eBook Electron Configurations Worksheet Answer Key
  - The Rise of Digital Reading Electron Configurations Worksheet Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configurations Worksheet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electron Configurations Worksheet Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configurations Worksheet Answer Key
  - Personalized Recommendations
  - Electron Configurations Worksheet Answer Key User Reviews and Ratings
  - Electron Configurations Worksheet Answer Key and Bestseller Lists

5. Accessing Electron Configurations Worksheet Answer Key Free and Paid eBooks
  - Electron Configurations Worksheet Answer Key Public Domain eBooks
  - Electron Configurations Worksheet Answer Key eBook Subscription Services
  - Electron Configurations Worksheet Answer Key Budget-Friendly Options
6. Navigating Electron Configurations Worksheet Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Electron Configurations Worksheet Answer Key Compatibility with Devices
  - Electron Configurations Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electron Configurations Worksheet Answer Key
  - Highlighting and Note-Taking Electron Configurations Worksheet Answer Key
  - Interactive Elements Electron Configurations Worksheet Answer Key
8. Staying Engaged with Electron Configurations Worksheet Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electron Configurations Worksheet Answer Key
9. Balancing eBooks and Physical Books Electron Configurations Worksheet Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electron Configurations Worksheet Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Electron Configurations Worksheet Answer Key
  - Setting Reading Goals Electron Configurations Worksheet Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configurations Worksheet Answer Key
  - Fact-Checking eBook Content of Electron Configurations Worksheet Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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