

Intellas

Job Automation for Payroll

Normal Biweekly Payroll Automation - Test Mode

Normal Biweekly Payroll Automation Automation

Type of Automation: **Normal Biweekly Automation**

Company: **ABC** **10/15/2021**

Start Date: **10/15/2021**

Last Start Date: **10/15/2021**

Next Start Date: **10/15/2021**

Number of Weeks: **1**

Last Payroll of the Month: **10/15/2021**

Run

Success: **0**

Cancel

Normal Biweekly Payroll Automation Automation

Type of Automation: **Normal Biweekly Automation**

Company: **ABC** **10/15/2021**

Start Date: **10/15/2021**

Last Start Date: **10/15/2021**

Next Start Date: **10/15/2021**

Number of Weeks: **1**

Last Payroll of the Month: **10/15/2021**

Run

Step	ID	Status	Result	Log Name
1	1001	Success	10/15/2021	10/15/2021
2	1002	Success	10/15/2021	10/15/2021
3	1003	Success	10/15/2021	10/15/2021
4	1004	Success	10/15/2021	10/15/2021
5	1005	Success	10/15/2021	10/15/2021
6	1006	Success	10/15/2021	10/15/2021

10/15/2021

Lawson Inventory Control User Guide

Chao Zhang



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 what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is
 a helpful way to remember this

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 oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ 24 3 $\text{Ag} + \text{Ni}^{3+} \rightarrow \text{Ni} + 3\text{Ag}$ oxidation reduction 25 $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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 magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a $\text{I}_2 + \text{O}_5$
 $5\text{Co} + \text{I}_2 \rightarrow 5\text{Co}^{2+} + 2\text{I}^-$ 5 $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$

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 reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation
 state and oxidation number concepts to identify redox reactions and to keep track of electrons

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 reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr^0 to Sr^{2+} oxidized reducing agent O_2 to O^{2-} reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li^0 to Li^{+1} oxidized red ag
 S^0 to S^{2-} reduced ox ag $3\text{C} + 2\text{Br}_2 \rightarrow \text{C}_2\text{Br}_4$

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