

Lab Report 2 – Characterization of Synthesis of Cadmium

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Matriculation No.: **21**

Course: **Chemie 1**

Reaction:

Reaction of pure cadmium metal with the H^+ of cadmium

Time (s)	Volume of H_2 (ml)	Pressure (atm)
0	0.00	0.00
10	0.00	0.00
20	0.00	0.00
Average H_2 (ml)		0.00
Average Pressure (atm)		0.00

Reaction: $\text{Pb} + \text{H}^+ \rightarrow \text{Pb}^{2+} + \text{H}_2$

a) Characterization of cadmium:

Weight of cadmium: **0.0000 g** Volume of H_2 : **0.0000 ml** Pressure: **0.0000 atm**

b) Cadmium reaction (synthesis):

Cadmium weight: **0.0000 g** Volume of H_2 : **0.0000 ml** Pressure: **0.0000 atm**

Reaction: $\text{Pb} + \text{H}^+ \rightarrow \text{Pb}^{2+} + \text{H}_2$

Volume of H_2 : **0.0000 ml** Pressure: **0.0000 atm**

Weight of cadmium: **0.0000 g** Volume of H_2 : **0.0000 ml** Pressure: **0.0000 atm**

Conclusion: Cadmium metal reacts with H^+ to form Pb^{2+} and H_2 . The reaction is exothermic and the H_2 gas is produced.

Reaction: $\text{Pb} + \text{H}^+ \rightarrow \text{Pb}^{2+} + \text{H}_2$. The reaction is exothermic and the H_2 gas is produced. The reaction is exothermic and the H_2 gas is produced. The reaction is exothermic and the H_2 gas is produced.

Notes: The reaction is exothermic and the H_2 gas is produced. The reaction is exothermic and the H_2 gas is produced. The reaction is exothermic and the H_2 gas is produced.

Gravimetric Determination Of Calcium Lab Report Answers

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