Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:

$$N_{2 (g)} + 3 H_{2 (g)} \leftrightarrow 2 NH_{3 (g)}$$

$$K = \frac{[NH_3]^2}{[N_2][H_2]^3}$$

$$I_{2(x)} + CI_{2(g)} \leftrightarrow 2ICI_{(g)}$$

$$K = \underline{\prod_{[Cl_2]}}$$

$$NO_{2(g)} \leftrightarrow NO_{(g)} + 1/2 O_{2(g)}$$

$$K = [NO][O_2]^{1/2}$$

 $[NO_2]$

 The dissociation of acetic acid, CH₃COOH, has an equilibrium constant at 25°C of 1.8 x 10⁻⁵. The reaction is

If the equilibrium concentration of CH_3COOH is 0.46 moles in 0.500 L of water and that of CH_3COO' is 8.1 x 10^{-3} moles in the same 0.500 L, calculate [H $^+$] for the reaction.

$$K = [CH_3COO^*][H^*]$$

$$[CH_3COOH]$$

Plug in known values and solve.

$$1.8 \times 10^{-5} = [8.1 \times 10^{-3} \text{ moles/0.500 L}] [H^*]$$

[0.46 moles/0.500 L]

$$[H^+] = 1.0 \times 10^{-3} M$$

Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

		catalyst	pressure	temperature	concentration
a.	speed	presence increase	direct relationship	direct relationship	direct relationship
ь.	pos"	nothing	based on moles	based on enthalpy of rxn	inverse relationship

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