

## **Matter And Energy Vocabulary**

<b>Word</b>	<b>Definition</b>
Alloy	A solution where two metals are dissolved into each other in the solid phase.
Amalgam	A solution where a metal is dissolved into mercury.
Aqueous	A solution where a solute is dissolved into water.
Atom	The smallest part of an element that still retains the properties of that element.
Calorimetry	The measurement of energy change between potential and kinetic energy by measuring the temperature change induced on a measured mass of water in a calorimeter.
Change	A transformation from one condition of matter to another.
Chemical Property	A property that can only be observed when a chemical change occurs.
Compound	Matter which results from the bonding of atoms of two or more elements to each other, decomposable into elements.
Element	Matter which exhibits definite physical and chemical properties unique to itself and different from all other forms of matter and cannot be decomposed into simpler forms of matter.
Endothermic	The conversion of kinetic energy into potential energy.
Exothermic	The conversion of potential energy into kinetic energy.
Heterogeneous	Matter that is unevenly distributed throughout a volume.
Homogeneous	Matter that is evenly distributed throughout a volume.
Ion	A charged particle formed when an atom gains or loses electrons.
Kinetic energy	Energy of motion.
Matter	That which exists with mass and volume.
Metal	An element that loses electrons when forming chemical bonds.
Metalloid	An element that exhibits properties of both metals and nonmetals.
Mixture	Matter of different types that are in physical proximity to each other, yet not chemically combined.
Nonmetal	An element that gains electrons from metals or shares electrons from other nonmetals when forming chemical bonds.
Periodic Table	A chart that lists the elements in order of increasing atomic number and arranges them in groups of similar chemical properties.
Physical Property	A change that can be observed without a chemical change occurring.
Potential energy	Stored energy, often stored in chemical bonds.
Property	A distinguishing characteristic of a sample of matter.
Solution	A homogeneous mixture consisting of a solute dissolved into a solvent.
Temperature	The average kinetic energy of a sample or system.
Tincture	A solution where a solute is dissolved into alcohol.

# Mark Rosengarten Answers Unit 2

**Michal Rosen-Zvi**



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