

Galvanic (Voltaic) Cells Problems

Name _____ Hr _____

1. Regarding the following reaction:



- List the species being oxidized: _____ List the species being reduced: _____
- Calculate E° for this cell.
- Which species receives electrons from the anode? _____
- Which species donates electrons to the cathode? _____

2. Given the following half-reactions:



- Write the overall equation for the galvanic cell. Be sure to balance the number of electrons.
- Calculate the E° for the cell.
- Designate which $\frac{1}{2}$ reaction occurs...
at the anode: _____ and at the cathode: _____
- Describe the direction of electron flow.

3. Find the following using a voltaic cell with the line notation: $\text{Al}_{(\text{s})}|\text{Al}^{3+}||\text{Pb}^{2+}|\text{Pb}_{(\text{s})}$

- Calculate the E° for the reaction.
- Write the balanced overall REDOX equation for the galvanic cell.

4. Using standard reduction potentials, calculate the cell potential (E°) for each of the following reactions:



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