

## 2.1.4 & 2.1.5 - Acids and redox

What are standard solutions?	A solution for which its concentration is known accurately.
What 3 things should you ensure when making standard solutions?	<ol style="list-style-type: none"> <li>1. Ensure you have clean dry equipment.</li> <li>2. Make washings to remove any salt <b>BEFORE MAKING</b> up the solution to the specified amount.</li> <li>3. Invert the volumetric flask to ensure everything is dissolved.</li> </ol>
How does ammonia react with acid?	<ul style="list-style-type: none"> <li>• ammonia + acid <math>\rightarrow</math> ammonium salt</li> <li>• <math>NH_3(aq) + HNO_3(aq) \rightarrow NH_4NO_3(aq)</math></li> </ul>
What are strong and weak acids ? Give 3 strong acids and 1 weak acid	<ul style="list-style-type: none"> <li>• Strong acids dissociate fully to release <math>H^+</math> ions in a solution e.g. <math>HCl \rightarrow H^+ + Cl^-</math></li> <li>• Weak acids dissociate partially to release <math>H^+</math> ions in a solution – exist in equilibrium</li> <li>• Strong acids include: <math>HCl</math>, <math>H_2SO_4</math>, and <math>HNO_3</math>.</li> <li>• Weak acids include carboxylic acids such as <math>CH_3COOH</math>.</li> </ul>
What are strong and weak bases? Give 1 weak base	<ul style="list-style-type: none"> <li>• Strong bases dissociate fully to release <math>OH^-</math> ions in a solution</li> <li>• Weak bases dissociate partially to release <math>OH^-</math> ions in a solution – exist in equilibrium e.g. <math>NH_3</math></li> </ul>
What are diprotic and triprotic acids (with examples) and what does this mean for neutralisation?	<ul style="list-style-type: none"> <li>• Diprotic acids and triprotic acids donate two and three protons respectively.</li> <li>• Examples include: <math>H_2SO_4</math> and <math>H_3PO_4</math>.</li> <li>• They need to react to the number of moles of base to be neutralised</li> </ul>

# Mark Scheme Ocr 2013 June Chemistry A172 C456

**AN Whitehead**



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