

7. A 645-gram beaker of water loses 12.5 kJ of water, if the water temperature of water started out at 49.5 °C, at what temperature did it end up?

a)  $\Delta T = q/mc$        $\Delta T = -12,500 \text{ J} / (45 \text{ g} \cdot 4.184 \text{ J/g}^\circ\text{C}) = -4.48^\circ\text{C}$

b)  $T_{\text{final}} = \Delta T + T_{\text{initial}}$        $T_{\text{final}} = -4.48^\circ\text{C} + 49.5^\circ\text{C} =$

**45.0 °C**

8. A 75.0 g piece of glass absorbs 6,350 joules of heat, if its initial temperature is 21.6 °C, what is its final temperature? Its specific heat capacity is 0.840 J/g °C.

a)  $\Delta T = q/mc$        $\Delta T = 6,350 \text{ J} / (75.0 \text{ g} \cdot 0.840 \text{ J/g}^\circ\text{C}) = 101^\circ\text{C}$

b)  $T_{\text{final}} = \Delta T + T_{\text{initial}}$        $T_{\text{final}} = 101^\circ\text{C} + 21.6^\circ\text{C} =$

**123 °C**

9. What is the specific heat capacity of mercury if 25.0 grams of it is heated from 25.0 °C to 155.0 °C while absorbing 455 joules of heat?

a)  $\Delta T = 155.0^\circ\text{C} - 25.0^\circ\text{C} = 130.0^\circ\text{C}$

b)  $c = q/m\Delta T$        $c = 455 \text{ J} / (25.0 \text{ g} \cdot 130.0^\circ\text{C}) =$

**0.140 J/gg °C**

10. How much heat is lost when 114 grams of lead cools from 155 °C to 22.2 °C? The specific heat capacity of lead is 0.129 J/g °C.

a)  $\Delta T = 22.2^\circ\text{C} - 155^\circ\text{C} = -133^\circ\text{C}$

b)  $q = (114 \text{ g})(0.129 \text{ J/g}^\circ\text{C})(-133^\circ\text{C}) =$

**1,960 J or 1.96 kJ**

# High School Chemistry Thermochemistry Questions

**Xiufeng Liu, William J. Boone**



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data fit and enhancement to Rasch models by Cognitive Diagnostic Models and Latent Class Analysis The volume provides readers who are new and experienced in applying Rasch measurement with advanced and exemplary applications in the forefront of various areas of science education research **Chemical News and Journal of Physical Science** ,1892

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## **Table of Contents High School Chemistry Thermochemistry Questions**

1. Understanding the eBook High School Chemistry Thermochemistry Questions
  - The Rise of Digital Reading High School Chemistry Thermochemistry Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying High School Chemistry Thermochemistry Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an High School Chemistry Thermochemistry Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from High School Chemistry Thermochemistry Questions
  - Personalized Recommendations
  - High School Chemistry Thermochemistry Questions User Reviews and Ratings

- High School Chemistry Thermochemistry Questions and Bestseller Lists
- 5. Accessing High School Chemistry Thermochemistry Questions Free and Paid eBooks
  - High School Chemistry Thermochemistry Questions Public Domain eBooks
  - High School Chemistry Thermochemistry Questions eBook Subscription Services
  - High School Chemistry Thermochemistry Questions Budget-Friendly Options
- 6. Navigating High School Chemistry Thermochemistry Questions eBook Formats
  - ePub, PDF, MOBI, and More
  - High School Chemistry Thermochemistry Questions Compatibility with Devices
  - High School Chemistry Thermochemistry Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of High School Chemistry Thermochemistry Questions
  - Highlighting and Note-Taking High School Chemistry Thermochemistry Questions
  - Interactive Elements High School Chemistry Thermochemistry Questions
- 8. Staying Engaged with High School Chemistry Thermochemistry Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers High School Chemistry Thermochemistry Questions
- 9. Balancing eBooks and Physical Books High School Chemistry Thermochemistry Questions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection High School Chemistry Thermochemistry Questions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine High School Chemistry Thermochemistry Questions
  - Setting Reading Goals High School Chemistry Thermochemistry Questions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of High School Chemistry Thermochemistry Questions
  - Fact-Checking eBook Content of High School Chemistry Thermochemistry Questions
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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