

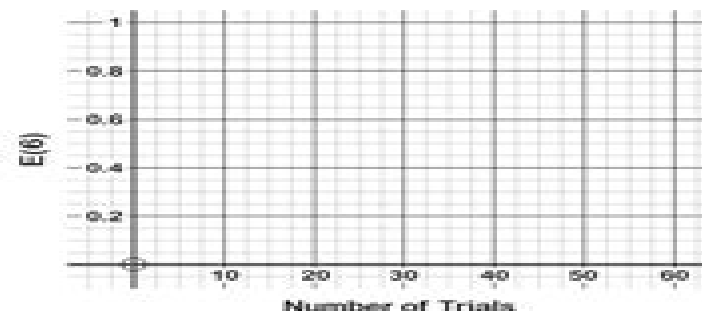
Experimental Probability



Section A

Roll a dice 60 times and record the number of '6's scored after every 10 rolls. Record your results in the table below. Leave your answers to two decimal places. Plot your results on the graph and then answer the questions.

Number of Trials	Number of '6's	Experimental probability of scoring '6'	Theoretical probability of scoring '6'
10			
20			
30			
40			
50			
60			

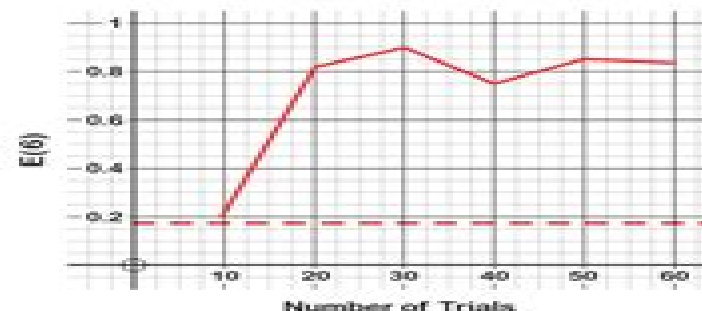


- 1) Were the results from your experiment as you expected? Why? Why not?
- 2) Which is the most reliable estimate of the experimental probability of scoring a '6'?
- 3) How could you get an experimental probability that is closer to the theoretical probability of scoring a '6'?

Section B

The same experiment is conducted using a different dice. The results are recorded in the table. Complete the table leaving answers to two decimal places. Plot the results on the graph and then answer the questions.

Number of Rolls	Number of '6's	Experimental probability of scoring '6'	Theoretical probability of scoring '6'
10	2		
20	16		
30	27		
40	30		
50	42		
60	50		



- 1) What does the horizontal line (in red) represent?
- 2) Is the dice a fair dice? How can you tell from your graph?
- 3) How would the graph compare with the horizontal line (theory line) as the number of trials increase?

Holt Mcdougal Worksheet Answers Theoretical And Experimental Probability

M Woodhall



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Holt Mcdougal Worksheet Answers Theoretical And Experimental Probability Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Holt Mcdougal Worksheet Answers Theoretical And Experimental Probability**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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web lewis structures represent the valence electrons present in the molecule as lone pairs unshared electrons drawn as dots or bonds shared electrons drawn as lines procedure for drawing lewis structures add up the total number of valence electrons that each atom contributes to the molecule polyatomic ion

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web the pcl 3 lewis structure has the typical case of phosphorus p in the center with 3 bonds to 3 other atoms phosphorus is from the same column as nitrogen in the periodic table meaning that p and n generally have the same bonding structure note the lone pair dots without bonds on top of p just like for n in the previous example for nh 3 chlorine cl

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web step 1 figure out how many electrons the molecule must have based on the number of valence electrons in each atom when drawing the structure of an ion be sure to add subtract electrons to account for the charge step 2 connect the atoms to each other with single bonds to form a skeleton structure

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web lewis structures a lewis structure is a representation of covalent molecules or polyatomic ions where all the valence electrons are shown distributed about the bonded atoms as either shared electron pairs bond pairs or unshared electron pairs lone pairs a shared pair of electrons is represented as a short line a single bond

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web jun 8 2022 lewis structure in chemistry lewis structures also called electron dot structures or electron dot diagrams are diagrams that show the bonding between atoms of a molecule and the lone pairs of

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web this exercise required us to take the formulas in the data table and calculate the number of valence electrons using the periodic table draw its lewis structure on paper build the molecular model and record a photo list the number of bonds and lone pairs surrounding the central atom and lastly identify the structure geometry