

- Directions:**
 1. An element may be assigned a positive or negative oxidation number.
 2. An atom with a higher electronegativity is assigned a negative oxidation number.
 3. The sum of all oxidation numbers for a molecule must be zero.

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1. Write an equation that shows how the element is oxidized or reduced.
 2. Write the oxidation number next to the element.
 3. Write the oxidation number next to the element.

Complete the following table, using the periodic table in the back of your book.

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	CHARGE OF ION	TYPE OF ION
1	Sodium	Na^+	11	10	+1	cation
2	Chlorine	Cl^-	17	18	-1	anion
3	Strontium	Sr^{2+}	38	36	+2	cation
4	Phosphorus	P^{3-}	15	18	-3	anion
5	Calcium	Ca^{2+}	20	18	+2	cation
6	Barium	Ba^{2+}	56	54	+2	cation
7	Hydrogen	H^+	1	0	+1	cation
8	Lithium	Li^+	3	2	+1	cation
9	Magnesium	Mg^{2+}	12	10	+2	cation
10	Sulfur	S^{2-}	16	18	-2	anion
11	Aluminum	Al^{3+}	13	10	+3	cation
12	Oxygen	O^{2-}	8	10	-2	anion
13	Iron	Fe^{2+}	26	24	+2	cation
14	Strontium	Sr^{2+}	38	36	+2	cation
15	Chlorine	Cl^-	17	18	-1	anion

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WJ Hussar

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located on the right side of the slide, partially overlapping the light blue bar.

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