

Model 2 – Prefixes and Suffixes

Prefix	Numerical Value
mono-	1
di-	2
tri-	3
tetra-	4
penta-	5
hexa-	6
hepta-	7
octa-	8
nona-	9
deca-	10

Molecular Formula	Name of Compound
BCl_3	Boron trichloride
SF_6	Sulfur hexafluoride
IF_7	Iodine heptafluoride
NI_3	Nitrogen triiodide
N_2O_4	Dinitrogen tetroxide
Cl_2O	Dichlorine monoxide
P_4O_{10}	Tetraphosphorus decoxide
B_5H_9	Pentaboron nonahydride
Br_2O_8	Tribromine octoxide
ClF	Chlorine monofluoride

- Examine the prefixes in Model 2. Fill in the numerical value that corresponds to each prefix.
- What suffix (ending) do all the compound names in Model 2 have in common?

They all end in "ide".

- Carefully examine the names of the compounds in Model 2. When is a prefix NOT used in front of the name of an element?

*When the first element is one. (Examples: IF_7 - iodine heptafluoride)
 NI_3 - nitrogen triiodide)*

- Consider the compound NO.

- Which element, nitrogen or oxygen, would require a prefix in the molecule name? Explain your answer.

Oxygen would require a prefix of "mono" because it is the second element and there is one oxygen. Nitrogen does not require a prefix because it is the first element and there is one nitrogen.

- Name the molecule NO.

*Nitrogen monoxide (Notice, the "o" of mono was dropped)
 ↳ not monoxide*

- Find two compounds in Model 2 that contain a subscript of "4" in their molecular formula.

- List the formulas and names for the two compounds.

N_2O_4 - dinitrogen tetroxide || P_4O_{10} - Tetraphosphorus decoxide

- What is different about the spelling of the prefix meaning "four" in these two names?

The "a" has been dropped from the end of the prefix tetra.

If8766 Answer Key Naming Molecular Compounds

Judd E. Hollander



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Table of Contents If8766 Answer Key Naming Molecular Compounds

1. Understanding the eBook If8766 Answer Key Naming Molecular Compounds
 - The Rise of Digital Reading If8766 Answer Key Naming Molecular Compounds
 - Advantages of eBooks Over Traditional Books
2. Identifying If8766 Answer Key Naming Molecular Compounds
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an If8766 Answer Key Naming Molecular Compounds
 - User-Friendly Interface
4. Exploring eBook Recommendations from If8766 Answer Key Naming Molecular Compounds
 - Personalized Recommendations
 - If8766 Answer Key Naming Molecular Compounds User Reviews and Ratings
 - If8766 Answer Key Naming Molecular Compounds and Bestseller Lists
5. Accessing If8766 Answer Key Naming Molecular Compounds Free and Paid eBooks

- If8766 Answer Key Naming Molecular Compounds Public Domain eBooks
- If8766 Answer Key Naming Molecular Compounds eBook Subscription Services
- If8766 Answer Key Naming Molecular Compounds Budget-Friendly Options
- 6. Navigating If8766 Answer Key Naming Molecular Compounds eBook Formats
 - ePub, PDF, MOBI, and More
 - If8766 Answer Key Naming Molecular Compounds Compatibility with Devices
 - If8766 Answer Key Naming Molecular Compounds Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of If8766 Answer Key Naming Molecular Compounds
 - Highlighting and Note-Taking If8766 Answer Key Naming Molecular Compounds
 - Interactive Elements If8766 Answer Key Naming Molecular Compounds
- 8. Staying Engaged with If8766 Answer Key Naming Molecular Compounds
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers If8766 Answer Key Naming Molecular Compounds
- 9. Balancing eBooks and Physical Books If8766 Answer Key Naming Molecular Compounds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection If8766 Answer Key Naming Molecular Compounds
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine If8766 Answer Key Naming Molecular Compounds
 - Setting Reading Goals If8766 Answer Key Naming Molecular Compounds
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of If8766 Answer Key Naming Molecular Compounds
 - Fact-Checking eBook Content of If8766 Answer Key Naming Molecular Compounds
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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