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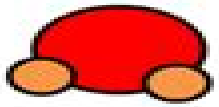
Chapter 6: Chemical Bonding Guided Notes

I. Why do Atoms Combine?

A. Chemical Formula

1. The chemical formula shows:
   a. The _____ in the compound.
   b. The ratio of their _____.



_____ ← H_2O → _____

B. Chemical Bond

1. Strong attractive force between _____ or ions in a molecule or compound.
2. Formed by:
   a. _____ electrons, (e-) by losing or gaining electrons.
   b. _____ electrons (e-)

C. Stability

1. Octet Rule
   a. Most atoms form bonds in order to have _____ valence electrons (e-).
   b. _____ outer energy level.
   c. Want to be like the Noble Gases (group 18).

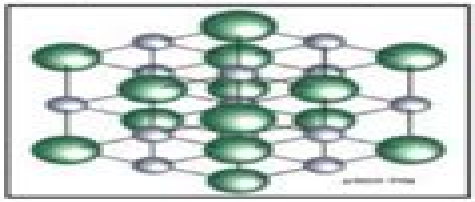
2. Stability is the driving force behind bond formation. All atoms bond to become stable.
3. Transferring e-

4. Sharing e-

II. Kinds of Chemical Bonds

A. Ionic Bond

1. Attraction between 2 _____ charged ions.
2. Ions are _____ atoms.
3. A _____ is a _____ charged ion. Forms when an atom loses an electron. Formed by _____.
4. A _____ is a _____ charged ion. Forms when an atom gains an electron. Formed by _____.
5. Formed by _____ electrons (e-) from a _____ to a _____.
6. Ions form a 3-D crystal lattice.



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Emilie Sanchez

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