

Accuracy, Precision & Errors

(Quia quiz: <http://www.quia.com/quiz/3507626.html>)

There is always an element of error in our measurements.

The **accuracy** of a measurement is how close it lies to the true value ("correct answer"). The **precision** of a measurement is how repeatable that measurement is.

Which of the dartboards shows:

- | | |
|--------------------------------------|-----------------------|
| 1. High precision and high accuracy? | Top - Middle - Bottom |
| 2. Low precision and low accuracy? | Top - Middle - Bottom |
| 3. High precision but low accuracy? | Top - Middle - Bottom |

Systematic error is an error that is the **same for all measurements** and can be adjusted. For example, a digital balance that is poorly calibrated might measure 0.5g too much each time.

Random error is **not consistent** between measurements and cannot be adjusted. For example, errors introduced by using different recording tools or people.

Which of the dartboards shows:

- | | | | |
|------------------|-----------------------|----------------------|-----------------------|
| 4. Random error? | Top - Middle - Bottom | 5. Systematic error? | Top - Middle - Bottom |
|------------------|-----------------------|----------------------|-----------------------|

Dartboard diagrams from http://preparatorychemistry.com/Bishop_Book_3_eBook.pdf



Measuring Temperature

Compare the glass thermometer with the temperature probe on the data logger.

Remember: Thermometers and dataloggers are NOT for stirring!

- ☐ Measure the temperature of the solution/ liquid – NOT the container
- ☐ Don't touch the thermometer or probe against the glass
- ☐ Wait for the temperature to stabilize before taking the reading

Glass thermometer °C (±)	
Temperature probe °C (±)	

Which would you choose to get quick readings of a reaction that changes temperature?

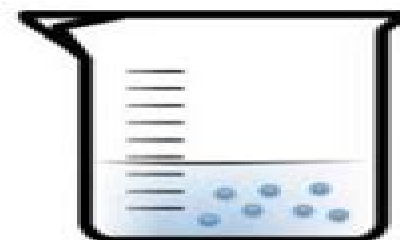
What is a Solution?

A solution is a **homogenous*** mixture of a substance in a liquid.

- The **solvent** is the liquid, e.g. water or alcohol, which dissolves the...
 - **Solute**, which is the solid, e.g. salt or sugar.
- ☐ Label the solvent and the solute in the diagram to the right.

*mixtures can be:

- **Homogenous** = evenly distributed, like solutions
- **Heterogeneous** = unevenly distributed, like soup or the oceans.



Beaker from: <http://www.wpclipart.com/science/beaker/beaker.png.html>

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