

Chapter 15 Equilibrium Problems

Section 15.1 - 15.2

1. Write the equilibrium expression for K_c for the following reactions:



2. Write the equilibrium constant expression for $\text{H}_2\text{(g)} + \text{Cl}_2\text{(g)} \rightleftharpoons 2\text{HCl(g)}$

3. Write the K_c expression for ammonia formation from nitrogen and hydrogen. K_c is 0.50 at 400°C.



Calculate K_c for this reaction with the decomposition

4. If at equilibrium $2\text{HCl(g)} \rightleftharpoons \text{H}_2\text{(g)} + \text{Cl}_2\text{(g)}$ at a temperature of 400°C, the mass of water at 0.01 M is 0.01. Calculate the value of K_c .

5. Write the equilibrium constant expressions for the acid-base pairs of the following reactions:



Equilibrium Ap Test Practice Problems Answers

Sabine Zange



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