

App 4: Thermodynamics – 12 August 2018, 12:00pm-12:30pm (with special discussion Tuesday 13:00-13:15, after lecture)

Remembering important facts about...

- Remembering **systems** **in equilibrium**
- Check the conditions: **isolated**, **isobaric**, **isochoric** for a thermodynamic system
- System has 2 given values: **temperature** and **volume**
- Change a **thermodynamic** **parameter** - remember that thermodynamic effects are **equilibrium** effects: **reversible** and **infinitesimal** **transformations**, **processes**, **phase changes**
- Change **state**, **parameters** for a **thermodynamic** **process**
- Calculate the **thermodynamic** **effect** given **state** **values**
- Remember that **entropy** **increases** **spontaneously** and **spontaneous** **not** **not**

Example: Thermodynamics of a piston with a gas (piston and gas together = System)

1. **State** of the system at the beginning and end of the process is known. At what is the thermodynamic parameter **temperature** kept? Is there a change in the volume of the system or volume **constant**?

```

p1 = 101325 Pa
V1 = 0.01 m^3
p2 = 202650 Pa
V2 = 0.02 m^3
n = 1
R = 8.314 J/(mol K)
T1 = p1 * V1 / (n * R)
T2 = p2 * V2 / (n * R)

```

2. **State** of the system at the beginning and end of the process

When the gas is at the initial state of the system, the pressure is 101325 Pa, the volume is 0.01 m³, the temperature is 273.15 K. When the gas is at the final state of the system, the pressure is 202650 Pa, the volume is 0.02 m³, the temperature is 546.3 K. The gas is at the initial state of the system, the pressure is 101325 Pa, the volume is 0.01 m³, the temperature is 273.15 K. The gas is at the final state of the system, the pressure is 202650 Pa, the volume is 0.02 m³, the temperature is 546.3 K.

```

p1 = 101325
V1 = 0.01
p2 = 202650
V2 = 0.02
n = 1
R = 8.314
T1 = p1 * V1 / (n * R)
T2 = p2 * V2 / (n * R)

```

3. The thermodynamic parameter **temperature** is kept at the initial state of the system, **isobaric**, by 273.15 K

```

p1 = 101325
V1 = 0.01
p2 = 101325
V2 = 0.02
n = 1
R = 8.314
T1 = p1 * V1 / (n * R)
T2 = p2 * V2 / (n * R)

```

4. The thermodynamic parameter **volume** is kept at the initial state of the system, **isochoric**, by 0.01 m³

When the gas is at the initial state of the system, the pressure is 101325 Pa, the volume is 0.01 m³, the temperature is 273.15 K. When the gas is at the final state of the system, the pressure is 202650 Pa, the volume is 0.01 m³, the temperature is 546.3 K. The gas is at the initial state of the system, the pressure is 101325 Pa, the volume is 0.01 m³, the temperature is 273.15 K. The gas is at the final state of the system, the pressure is 202650 Pa, the volume is 0.01 m³, the temperature is 546.3 K.

```

p1 = 101325
V1 = 0.01
p2 = 202650
V2 = 0.01
n = 1
R = 8.314
T1 = p1 * V1 / (n * R)
T2 = p2 * V2 / (n * R)

```

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The Complete Chemistry Study Guide and Note Cards and MCAT Konstantinos Papadopoulos, 2012-07-06

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