

## Dynamic Equilibrium

At what point is a reversible reaction “completed?”

### Why?

Most of the reactions we have studied year have been “forward” reactions—once the reactant has been changed into the product it stays that way. The reaction “stops” when the limiting reactant runs out.

However, in a reversible reaction (one that takes place in both directions), the product can turn back into reactant in a continuous cycle that never stops. So...

When can we say the reaction is “finished?”

Can we predict how much reactant and how much product will be present when the reaction is “done?”

### Model 1 – A Reversible Reaction



1. What is the significance of the double arrow in Model 1?
2. Imagine the reaction in Model 1 starts with 100 molecules of A and zero molecules of B in a 1 liter container. Explain why the concentration of reactant A will never reach zero.
3. Imagine the reaction in Model 1 starts with zero molecules of A and 100 molecules of B in a 1 liter container. Will the concentration of product B increase, decrease, or remain the same as the reaction proceeds? Explain.

# Equilibrium A Reversible Reaction Pogil Answers

**Philip George Ashmore**



## Equilibrium A Reversible Reaction Pogil Answers:

**Science Citation Index**, 1992 Vols for 1964 have guides and journal lists      **A Study of the Equilibrium of the Reaction  $3\text{K}_2\text{MnO}_4 + 2\text{H}_2\text{O}$  [reversible Reaction]  $2\text{KMnO}_4 + 4\text{KOH}$**  Chester E. Linscott, 1918      **Reversible Reactions and Dynamic Equilibrium : Utilization Guide** Jarvis, Gerry, Alberta Educational Communications Corporation, 1993      *The Nature of Equilibrium* Lifeliqe, 2019 This lesson plan covers the nature of reversible reactions chemical equilibrium and how to calculate and use equilibrium constants      *Principles of Chemical Equilibrium* Philip George Ashmore, 1961      **Equilibrium Constant and Rates for the Reversible Reaction  $\text{N}_4(+)$**  R. N. Varney, LOCKHEED MISSILES AND SPACE CO PALO ALTO CALIF LOCKHEED PALO ALTO RESEARCH LAB., 1968 The equilibrium currents of  $\text{N}_4$  and  $\text{N}_2$  in the reaction  $\text{N}_4 \rightleftharpoons 2\text{N}_2$  were measured as a function of  $E/p$  the ratio of electric field strength to pressure a quantity which is convertible to equivalent ion temperature The range of  $E/p$  is from 40 to 70 V cm Torr a range in which the equilibrium of the reaction shifts strongly from the left toward the right Both the forward and the backward rate constants were also measured Values are given as a function of  $E/p$  The range of  $p$  is from 0.63 to 1.24 Torr but comparisons are also made with older data at 15 to 35 Torr The ratio of currents of  $\text{N}_4$  and  $\text{N}_2$  is shown to depend on pressure as predicted from earlier thermodynamic studies Author      **Chemical Equilibrium** Robert Blanchard Fischer, Dennis G. Peters, 1970      **Principles of Chemical Equilibrium** Kelso B. Morris, 1965      Equilibrium Calculations John T. Donoghue, 1971      **Chemical Equilibria** Philip J. Hills, 1969      **Chemical Equilibrium** Allen J. Bard, 1966      **The Non-equilibrium Kinetics of Reversible Bimolecular Reactions**, 1999      *Workbook 15* Thobeka Pete, Turret Correspondence College (Johannesburg), 1988      **Chemical Reactions 2 : Equilibrium and Oxidation-reduction : CHE-5043-2 : Learning Guide** Chunn, Barbara, Dumas, André, Céline Tremblay, Moisan, Mireille, Société de formation à distance des commissions scolaires du Québec, Morazain, Paule, De Fulviis, Claudia, 2001      **Equilibrium III**, 1984      Chemical Equilibrium William B. Guenther, 1975      *Chemical equilibrium*,      Equilibrium Constants of Liquid-liquid Distribution Reactions International union of pure and applied chemistry. Commission on equilibrium data,

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