

1. According to the Brønsted—Lowry definition, which chemical species can function *both* as an acid and as a base?

- $\text{Cl}^-$
- $\text{SO}_4^{2-}$
- $\text{NH}_4^+$
- $\text{HCO}_3^-$
- $\text{H}_3\text{O}^+$

2. In the reaction  $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$  which is an acid—base conjugate pair?

- $\text{H}_2\text{O}$  and  $\text{HCN}$
- $\text{H}_2\text{O}$  and  $\text{OH}^-$
- $\text{CN}^-$  and  $\text{H}_2\text{O}$
- $\text{HCN}$  and  $\text{OH}^-$

3. Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX

- a. It is less basic than a 1 M solution of NaY.
- b. It is more basic than a 1 M solution of NaY.
- c. It yields a neutral solution.
- d. It is more concentrated than a 1 M solution of NaY.

4. The pH of a 0.03 M HCl solution is

a.	1.5
b.	2.5
c.	3.5
d.	12.5

5. The pH of a  $1.0 \times 10^{-3}$  M  $\text{Ba}(\text{OH})_2$  solution at  $25^\circ\text{C}$  is

a.	2.7
b.	3.0
c.	11.0
d.	11.3

6. Which is the strongest acid?

- $\text{HClO}$
- $\text{HClO}_2$
- $\text{HClO}_3$
- $\text{HClO}_4$

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

a.  $\text{NaC}_2\text{H}_3\text{O}_2$   
b.  $\text{NaNO}_3$   
c.  $\text{NH}_4\text{Cl}$   
d.  $\text{BaSO}_4$

8. The weakest of the bases listed is

Acid	Conjugate Base	K <sub>a</sub> (Ionization Constant of Acid)
HCl	Cl <sup>-</sup>	100% ionized
HSO <sub>4</sub> <sup>-</sup>	SO <sub>4</sub> <sup>2-</sup>	1.2 x 10 <sup>-2</sup>
H <sub>2</sub> S	HS <sup>-</sup>	5.7 x 10 <sup>-8</sup>
HS <sup>-</sup>	S <sup>2-</sup>	1.2 x 10 <sup>-14</sup>

- $\text{Cl}^-$
- $\text{CN}^-$
- $\text{HS}^-$
- $\text{S}^{2-}$
- $\text{SO}_3^{2-}$

Year	Atlantic croaker	Atlantic menhaden	Atlantic herring	Atlantic bluefish
1960	45	45	45	45
1965	40	40	40	40
1970	35	35	35	35
1975	30	30	30	30
1980	25	25	25	25
1985	20	20	20	20
1990	15	10	10	10

10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate  $\text{KC}_2\text{H}_3\text{O}_2$  is less alkaline than a 0.10 M solution of potassium cyanide, KCN?

- Hydrocyanic acid is a weaker acid than acetic acid.
- Hydrocyanic acid is less soluble in water than acetic acid.
- Cyanides are less soluble than acetates.
- Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of

- 0.100 M sodium solution,
- 0.0500 M sodium hydroxide solution,
- 0.0500 M benzoic acid solution,
- 0.0500 M sodium benzoate solution.

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**Oscar Krisen Buros**



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