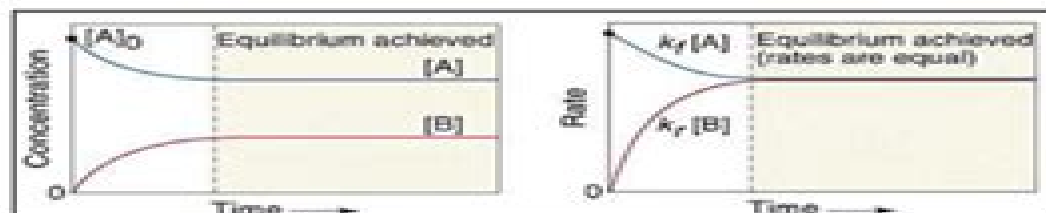


## Chapter 14 Worksheet 1 (ws14.1)

### Chemical Equilibrium, the Equilibrium Constant ( $K_{eq}$ ), and the Reaction Quotient ( $Q$ )

As a reaction proceeds in a **closed system**, the concentrations of reactants and products change until, at **equilibrium**, the rate of the forward reaction is equal to the rate of the reverse reaction. At that point there is **no net change** in the concentrations of reactants or products. For example, for the simple reaction:  $A \rightleftharpoons B$ :



**Regardless of the initial concentrations**, the concentrations of reactants and products present **at equilibrium** are related through the equilibrium constant ( $K_{eq}$ ) as follows:

For the reaction:  $aA + bB \rightleftharpoons cC + dD$

If all reactants and products are in solution then:

$$K_{eq} = K_c = \frac{[C]_{eq}^c [D]_{eq}^d}{[A]_{eq}^a [B]_{eq}^b}$$

(The subscript “eq” indicates “equilibrium” concentration. The subscript “c” stands for “concentration”. The exponents are the coefficients from the balanced equation.)

If all reactants and products are ideal gases then **either** partial pressures or concentrations can be used in the equilibrium constant expression. In terms of partial pressures:

$$K_{eq} = K_p = \frac{(P_C)_{eq}^c (P_D)_{eq}^d}{(P_A)_{eq}^a (P_B)_{eq}^b} \quad \text{(The subscript “p” stands for “pressure”).}$$

The numerical value of the two equilibrium constants above will be different. Remember the ideal gas law shows that the partial pressure of a gas is directly proportional to its molarity:

$$P = \frac{nRT}{V} = MRT \quad (M \text{ is molarity and } R = 0.0821 \text{ L-atm/mol-K}).$$

Using this relationship, it is easy to show that (see pages 605-606 in your textbook) :

$$K_p = K_c(RT)^{\Delta n} \quad \text{where } \Delta n = \text{moles of gaseous products} - \text{moles of gaseous reactants}$$

For reactions involving both gases and aqueous solutions, molar concentrations **and** partial pressures can appear in the equilibrium constant expression.

# Equilibrium Ice Practice Problems

**CH Cherryholmes**



## **Equilibrium Ice Practice Problems:**

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Untersteiner, 2013-06-29 The Advanced Study Institute Ice Physics in the Natural and Endangered Environment was held at Acquafredda di Maratea Italy from September 7 to 19 1997 The ASI was designed to study the broad range of ice science and technology and it brought together an appropriately interdisciplinary group of lecturers and students to study the many facets of the subject The talks and poster presentations explored how basic molecular physics of ice have important environmental consequences and conversely how natural phenomena present new questions for fundamental study The of lectures discusses these linkages in order that overall unity of following summary the subject and this volume can be perceived Not all of the lecturers and participants were able to contribute a written piece but their active involvement was crucial to the success of the Institute and thereby influenced the content of the volume We began the Institute by retracing the history of the search for a microscopic understanding of melting Our motivation was straightforward Nearly every phenomenon involving ice in the environment is influenced by the change of phase from solid to liquid or vice versa Hence a sufficiently deep physical picture of the melting transition enriches our appreciation of a vast array of geophysical and technical problems *Bihar STET Paper II : Chemistry 2024 (English Edition) | Higher Secondary (Class 11 & 12) - Bihar*

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