

Intermolecular Force Worksheet Key

1. Draw the following substances. Then, identify **the strongest** intermolecular force present in pure samples of the following substances:



dipole-dipole forces



hydrogen bonds



dipole-dipole forces



dipole-dipole forces



dipole-dipole forces



London dispersion forces

2. Identify the strongest intermolecular force operating in the condensed phases of the following substances. Fully explain how you determined this.

a. Cl₂ London dispersion forces

The Cl-Cl bond is nonpolar so the molecule is nonpolar. Non polar molecules have only London dispersion forces operating in the substance.

b. CO Dipole-dipole forces

The C-O bond is polar so the molecule is polar. Polar molecules have dipole-dipole forces. They also have London dispersion forces, but dipole-dipole forces are stronger.

c. SO₂ Dipole-dipole forces

SO₂ is a bent, polar molecule. The strongest intermolecular force in a polar molecule is the dipole-dipole force

d. CH₂Cl₂ Dipole-dipole forces

The strongest intermolecular force in a polar molecule that cannot form hydrogen bonds is the dipole-dipole force

e. HF Hydrogen bonding forces

Molecules that have hydrogen attached to an O, N, or F can form hydrogen bonds. These are the strongest of the intermolecular forces.

g. CH₃-O-CH₃ Dipole-dipole forces

The hydrogen atoms are not bonded to the oxygen, so this molecule cannot form hydrogen bonds. It is polar, so it will have dipole-dipole forces.

Intermolecular Forces And Organics Worksheet Answer Key

S Ben Porath



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