

Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl . In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

1	H ⁺												
2	Li ⁺	Be ²⁺	Transition elements						N ³⁻	O ²⁻	F ⁻		
3	Na ⁺	Mg ²⁺					Al ³⁺		P ³⁻	S ²⁻	Cl ⁻		
4	K ⁺	Ca ²⁺	Fe ²⁺ Fe ³⁺	Ni ²⁺ Ni ³⁺	Cu ⁺ Cu ²⁺	Zn ²⁺						Br ⁻	
5	Rb ⁺	Sr ²⁺			Ag ¹⁺			Sn ²⁺ Sn ⁴⁺				I ⁻	
6		Ba ²⁺				Hg ₂ ²⁺ Hg ²⁺		Pb ²⁺ Pb ⁴⁺					
	Cations						Anions						

1. Based on the information in Model 1:

- a. Identify three elements that form only one cation. $\leftarrow$ Positive Ion
 Na^+ , Sr^{2+} , Al^{3+} (Any box with one element works.)
- b. Identify three elements that form only one anion. $\leftarrow$ Negative Ion
 P^{3-} , O^{2-} , Br^- (Any box with one element works.)
- c. Identify three elements that form more than one cation.
 Fe^{+2} and Fe^{+3} , Ni^{+2} and Ni^{+3} , Cu^{+1} and Cu^{+2}
- d. In what region of the periodic table are these "multiple ion" elements usually located?
Transition elements or "d" block

2. Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.
- $2K^+$ and S^{2-} or K_2S
3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



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Table of Contents Naming Ionic Compounds Packet Answers Pogill

1. Understanding the eBook Naming Ionic Compounds Packet Answers Pogill
 - The Rise of Digital Reading Naming Ionic Compounds Packet Answers Pogill
 - Advantages of eBooks Over Traditional Books
2. Identifying Naming Ionic Compounds Packet Answers Pogill
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Naming Ionic Compounds Packet Answers Pogill
 - User-Friendly Interface
4. Exploring eBook Recommendations from Naming Ionic Compounds Packet Answers Pogill
 - Personalized Recommendations
 - Naming Ionic Compounds Packet Answers Pogill User Reviews and Ratings

- Naming Ionic Compounds Packet Answers Pogill and Bestseller Lists
- 5. Accessing Naming Ionic Compounds Packet Answers Pogill Free and Paid eBooks
 - Naming Ionic Compounds Packet Answers Pogill Public Domain eBooks
 - Naming Ionic Compounds Packet Answers Pogill eBook Subscription Services
 - Naming Ionic Compounds Packet Answers Pogill Budget-Friendly Options
- 6. Navigating Naming Ionic Compounds Packet Answers Pogill eBook Formats
 - ePub, PDF, MOBI, and More
 - Naming Ionic Compounds Packet Answers Pogill Compatibility with Devices
 - Naming Ionic Compounds Packet Answers Pogill Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Naming Ionic Compounds Packet Answers Pogill
 - Highlighting and Note-Taking Naming Ionic Compounds Packet Answers Pogill
 - Interactive Elements Naming Ionic Compounds Packet Answers Pogill
- 8. Staying Engaged with Naming Ionic Compounds Packet Answers Pogill
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Naming Ionic Compounds Packet Answers Pogill
- 9. Balancing eBooks and Physical Books Naming Ionic Compounds Packet Answers Pogill
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Naming Ionic Compounds Packet Answers Pogill
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Naming Ionic Compounds Packet Answers Pogill
 - Setting Reading Goals Naming Ionic Compounds Packet Answers Pogill
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Naming Ionic Compounds Packet Answers Pogill
 - Fact-Checking eBook Content of Naming Ionic Compounds Packet Answers Pogill
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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