

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has. solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1802	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

Final Exam Study Guide Physical Science

SJ Ball



Final Exam Study Guide Physical Science:

Embracing the Beat of Appearance: An Emotional Symphony within **Final Exam Study Guide Physical Science**

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Table of Contents Final Exam Study Guide Physical Science

1. Understanding the eBook Final Exam Study Guide Physical Science
 - The Rise of Digital Reading Final Exam Study Guide Physical Science
 - Advantages of eBooks Over Traditional Books
2. Identifying Final Exam Study Guide Physical Science
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Final Exam Study Guide Physical Science
 - User-Friendly Interface
4. Exploring eBook Recommendations from Final Exam Study Guide Physical Science
 - Personalized Recommendations

- Final Exam Study Guide Physical Science User Reviews and Ratings
- Final Exam Study Guide Physical Science and Bestseller Lists
- 5. Accessing Final Exam Study Guide Physical Science Free and Paid eBooks
 - Final Exam Study Guide Physical Science Public Domain eBooks
 - Final Exam Study Guide Physical Science eBook Subscription Services
 - Final Exam Study Guide Physical Science Budget-Friendly Options
- 6. Navigating Final Exam Study Guide Physical Science eBook Formats
 - ePub, PDF, MOBI, and More
 - Final Exam Study Guide Physical Science Compatibility with Devices
 - Final Exam Study Guide Physical Science Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Final Exam Study Guide Physical Science
 - Highlighting and Note-Taking Final Exam Study Guide Physical Science
 - Interactive Elements Final Exam Study Guide Physical Science
- 8. Staying Engaged with Final Exam Study Guide Physical Science
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Final Exam Study Guide Physical Science
- 9. Balancing eBooks and Physical Books Final Exam Study Guide Physical Science
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Final Exam Study Guide Physical Science
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Final Exam Study Guide Physical Science
 - Setting Reading Goals Final Exam Study Guide Physical Science
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Final Exam Study Guide Physical Science
 - Fact-Checking eBook Content of Final Exam Study Guide Physical Science

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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