Skill - tire owership to calculate the missing quantity given the others.

Be sure to show all work required for the following calculations:

- A 15.75g prece of iron absorbs 1986.75) of heat and its temperature changes from 25 to 175°C. Calculate the specific heat canacity of iron.
 - 10. How many joules of heat are needed to raise the temp of 10.0g of aluminum from 22 to 55°C. If the specific heat of aluminum is 0.90 µg°C?
 - 11. To what temperature will a 50 Sq piece of glass raise if it absorbs \$275 j of heat and its specific heat capacity is \$50 Vg*C? The initial temperature of the glass is 26°C.
 - 12. Calculate the specific heat capacity of a piece of wood if 1500 0g of the wood absorbs 6.75×10%, and its temp changes from 32 to 57°C.
 - 13. 100mL of 4.0°C water is heated until its temp is 37°C. If the specific heat capacity of water is 4.10 Mg°C, calculate the amount of heat energy needed to cause this rise in temp in joules AND calories.
 - 14. 25. bg of an unknown substance is heated from 25 to 155°C, absorbing 455§ in the process Calculate the specific heat capacity of the unknown substance. Based on your answer determine the identity of the substance, using the chart on a 229 of your text.
- Skill . Calculate the enthaloy associated with a reaction.
 - 15. Using the following balanced equation, calculate the thermal energy released (5H) when 45 Dg of toward, dissolved in water?

HH JOSO - HH / (ag) + NO, (ag)

AM = 200.5 M

- **16.** Calculate the change in enthalpy LDM when 1.00g of hydrogen peroxide decomposes. $H_{\nu}O_{\nu}(0) + H_{\nu}O(0) + H_{\nu}O_{\nu}(0)$ $\Delta H = -90.2 \text{ kg}$
- **SMB** Calculate the energy of a reaction using Hearts Law and identify if it is existly entitle exist.

Entropy Section Review Answers

Clemens Wendtner



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