

Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion (-)	Polyatomic Anion Name
HClO_3	Chloric acid	H_3O^+	ClO_3^{-1}	Chlorate
H_2SO_3	Sulfurous acid	$2\text{H}_3\text{O}^+$	SO_3^{-2}	Sulfite
H_2SO_4	Sulfuric acid	$2\text{H}_3\text{O}^+$	SO_4^{-2}	Sulfate
H_3PO_3	Phosphorous acid	$3\text{H}_3\text{O}^+$	PO_3^{-3}	Phosphite
H_3PO_4	Phosphoric acid	$3\text{H}_3\text{O}^+$	PO_4^{-3}	Phosphate
HNO_3	Nitric acid	H_3O^+	NO_3^{-1}	Nitrate
HNO_2	Nitrous acid	H_3O^+	NO_2^{-1}	Nitrite
H_2CO_3	Carbonic acid	$2\text{H}_3\text{O}^+$	CO_3^{-2}	Carbonate

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

Ternary acids contain three elements and they also contain the element oxygen.

b. What number do you think the prefix "ter-" refers to?

"ter" refers to three

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming HClO_3 ?

ClO_3^{-1} has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCl , not HClO_3 .

10. Write a rule for naming ternary acids.

- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.

- "ate" is replaced with "ic" and "ite" is replaced with "ous".

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



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N Colangelo



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The Particulate Nature of Polyatomic Ions, 2007 **Investigations Into the Origins of Polyatomic Ions in Inductively Coupled Plasma-mass Spectrometry** Sally M. McIntyre, 2010

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Table of Contents Naming Polyatomic Ions Packet Answers Pogill

1. Understanding the eBook Naming Polyatomic Ions Packet Answers Pogill
 - The Rise of Digital Reading Naming Polyatomic Ions Packet Answers Pogill
 - Advantages of eBooks Over Traditional Books
2. Identifying Naming Polyatomic Ions Packet Answers Pogill
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Naming Polyatomic Ions Packet Answers Pogill
 - User-Friendly Interface
4. Exploring eBook Recommendations from Naming Polyatomic Ions Packet Answers Pogill
 - Personalized Recommendations
 - Naming Polyatomic Ions Packet Answers Pogill User Reviews and Ratings
 - Naming Polyatomic Ions Packet Answers Pogill and Bestseller Lists
5. Accessing Naming Polyatomic Ions Packet Answers Pogill Free and Paid eBooks
 - Naming Polyatomic Ions Packet Answers Pogill Public Domain eBooks
 - Naming Polyatomic Ions Packet Answers Pogill eBook Subscription Services
 - Naming Polyatomic Ions Packet Answers Pogill Budget-Friendly Options

6. Navigating Naming Polyatomic Ions Packet Answers Pogill eBook Formats
 - ePub, PDF, MOBI, and More
 - Naming Polyatomic Ions Packet Answers Pogill Compatibility with Devices
 - Naming Polyatomic Ions Packet Answers Pogill Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Naming Polyatomic Ions Packet Answers Pogill
 - Highlighting and Note-Taking Naming Polyatomic Ions Packet Answers Pogill
 - Interactive Elements Naming Polyatomic Ions Packet Answers Pogill
8. Staying Engaged with Naming Polyatomic Ions Packet Answers Pogill
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Naming Polyatomic Ions Packet Answers Pogill
9. Balancing eBooks and Physical Books Naming Polyatomic Ions Packet Answers Pogill
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Naming Polyatomic Ions Packet Answers Pogill
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Naming Polyatomic Ions Packet Answers Pogill
 - Setting Reading Goals Naming Polyatomic Ions Packet Answers Pogill
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Naming Polyatomic Ions Packet Answers Pogill
 - Fact-Checking eBook Content of Naming Polyatomic Ions Packet Answers Pogill
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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