

## Assignment #2 – Compound Names and Formulas

### Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide Cu<sub>2</sub>O
2. mercury (I) oxide Hg<sub>2</sub>O
3. gold (III) chloride AuCl<sub>3</sub>
4. thallium (III) bromide TlBr<sub>3</sub>
5. bismuth (V) oxide Bi<sub>2</sub>O<sub>5</sub>
6. terbium (III) oxide Tb<sub>2</sub>O<sub>3</sub>
7. uranium (VI) oxide UO<sub>3</sub>
8. protactinium (IV) bromide PaBr<sub>4</sub>
9. cerium (III) oxide Ce<sub>2</sub>O<sub>3</sub>
10. arsenic (V) sulphide As<sub>2</sub>S<sub>5</sub>

11. manganese (III) oxide Mn<sub>2</sub>O<sub>3</sub>
12. vanadium (II) bromide VBr<sub>2</sub>
13. niobium(V) oxide Nb<sub>2</sub>O<sub>5</sub>
14. titanium (III) oxide Ti<sub>2</sub>O<sub>3</sub>
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co<sub>3</sub>P<sub>2</sub>
18. tin (II) oxide SnO
19. thulium (II) bromide TmBr<sub>2</sub>
20. copper (II) bromide CuBr<sub>2</sub>

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. SnCl<sub>4</sub> Tin (IV) Chloride
2. BiBr<sub>5</sub> Bismuth (V) Bromide
3. PoO<sub>2</sub> Polonium (IV) Oxide
4. PbI<sub>2</sub> Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au<sub>2</sub>O<sub>3</sub> Gold (III) Oxide
8. FeCl<sub>2</sub> Iron (II) Chloride

9. PdF<sub>4</sub> Palladium (IV) Fluoride
10. Os<sub>2</sub>O<sub>3</sub> Osmium (III) Oxide
11. MoBr<sub>2</sub> Molybdenum (II) Bromide
12. VCl<sub>5</sub> Vanadium (V) Chloride
13. Mn<sub>2</sub>O<sub>3</sub> Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np<sub>2</sub>O<sub>3</sub> Neptunium (III) Oxide
16. V<sub>2</sub>O<sub>5</sub> Vanadium (V) Oxide

# Names And Formulas Worksheet 1 Answers

**M Carnoy**



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### **Table of Contents Names And Formulas Worksheet 1 Answers**

1. Understanding the eBook Names And Formulas Worksheet 1 Answers
  - The Rise of Digital Reading Names And Formulas Worksheet 1 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Names And Formulas Worksheet 1 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Names And Formulas Worksheet 1 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Names And Formulas Worksheet 1 Answers
  - Personalized Recommendations

- Names And Formulas Worksheet 1 Answers User Reviews and Ratings
- Names And Formulas Worksheet 1 Answers and Bestseller Lists
- 5. Accessing Names And Formulas Worksheet 1 Answers Free and Paid eBooks
  - Names And Formulas Worksheet 1 Answers Public Domain eBooks
  - Names And Formulas Worksheet 1 Answers eBook Subscription Services
  - Names And Formulas Worksheet 1 Answers Budget-Friendly Options
- 6. Navigating Names And Formulas Worksheet 1 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Names And Formulas Worksheet 1 Answers Compatibility with Devices
  - Names And Formulas Worksheet 1 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Names And Formulas Worksheet 1 Answers
  - Highlighting and Note-Taking Names And Formulas Worksheet 1 Answers
  - Interactive Elements Names And Formulas Worksheet 1 Answers
- 8. Staying Engaged with Names And Formulas Worksheet 1 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Names And Formulas Worksheet 1 Answers
- 9. Balancing eBooks and Physical Books Names And Formulas Worksheet 1 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Names And Formulas Worksheet 1 Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Names And Formulas Worksheet 1 Answers
  - Setting Reading Goals Names And Formulas Worksheet 1 Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Names And Formulas Worksheet 1 Answers
  - Fact-Checking eBook Content of Names And Formulas Worksheet 1 Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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