

1. Consider the following reaction:



An equilibrium mixture of this reaction at a certain temp was found to have $[\text{CO}] = 0.115\text{M}$, $[\text{H}_2] = 0.110\text{M}$, and $[\text{CH}_3\text{OH}] = 0.190\text{M}$. What is the value of the equilibrium constant (K_c) at this temperature?

$$K_c = \frac{[\text{CH}_3\text{OH}]}{[\text{CO}][\text{H}_2]^2} = \frac{(0.19)}{(0.115)(0.11)^2} = 136.5$$

2. Consider the following reaction:



An equilibrium mixture of this reaction at a certain temperature was found to have $[\text{NH}_3] = 0.274\text{M}$ and $[\text{H}_2\text{S}] = 0.355\text{M}$. What is the value of the equilibrium constant (K_c) at this temperature?

$$K_c = [\text{NH}_3][\text{H}_2\text{S}] = (0.274)(0.355) = 0.0973$$

3. Consider the following reaction:



The reaction between CO and H_2 is carried out at a specific temperature with initial concentrations of $\text{CO} = 0.32\text{M}$ and $\text{H}_2 = 0.45\text{M}$. At equilibrium, the concentration of CH_3OH is 0.14M . Find the equilibrium constant at this temperature.

	CO	2H_2	CH_3OH
I	0.32	0.45	
C	-x	-2x	+x
E	0.18	0.17	0.14

$$K_c = \frac{(0.14)}{(0.18)(0.17)^2} = 26.9$$

4. Consider the following reaction:



The reaction of CH_4 is carried out at some temperature with an initial concentration of $[\text{CH}_4] = 0.090\text{M}$. At equilibrium, the concentration of H_2 is 0.012M . Find the equilibrium constant at this temperature.

	2CH_4	C_2H_2	3H_2
I	0.09		
C	-2x	+x	+3x
E	0.082	0.004	0.012M

$$K_c = \frac{(0.004)(0.012)^3}{(0.082)^2} = 1.0 \times 10^{-6}$$

5. Consider the following reaction:



A reaction mixture contains $[\text{NO}_2] = 0.0248\text{M}$ and $[\text{N}_2\text{O}_4] = 0.0337\text{M}$. Calculate Q_c at the given concentrations.

$$Q_c = \frac{(0.0248)^2}{(0.0337)} = 0.0183 \quad Q_c > K_c, \text{ rxn goes left}$$

6. The following reaction was carried out in a 2.75L reaction vessel at 1100K :



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