

## Balancing Nuclear Equations

1. Write balanced equations for the following nuclear reactions:

- (A) The isotope C-11 decays by electron capture.
- (B) Curium-246 reacts with carbon-12 to produce nobelium-254 and four neutrons.
- (C) Californium-250 reacts with boron-10 to produce lawrencium-258 and two neutrons.
- (D) Fluorine-18 decays to oxygen-18 by positron emission.
- (E) Sodium-24 decays by beta emission.
- (F) Krypton-76 absorbs a beta particle to form bromine-76.

2. Fill in the blanks to complete the following nuclear reactions.

- (A)  ${}^1\text{H}_1 + {}^7\text{Li}_3 \rightarrow$  \_\_\_\_\_
- (B)  ${}^9\text{Be}_4 + {}^1\text{n}_0 \rightarrow {}^4\text{He}_2 +$  \_\_\_\_\_  $+ {}^4\text{He}_2$
- (C)  ${}^{27}\text{Al}_{13} + {}^4\text{He}_2 \rightarrow$  \_\_\_\_\_  $+ {}^{30}\text{P}_{15}$
- (D)  ${}^{63}\text{Cu}_{29} + {}^2\text{H}_1 \rightarrow 2{}^1\text{n}_0 +$  \_\_\_\_\_
- (E)  ${}^9\text{Be}_4 + {}^4\text{He}_2 \rightarrow$  \_\_\_\_\_  $+ {}^{12}\text{C}_6$
- (F)  ${}^{239}\text{Np}_{93} \rightarrow {}^0\text{e}_{-1} +$  \_\_\_\_\_
- (G)  ${}^4\text{He}_2 + {}^{238}\text{U}_{92} \rightarrow 2{}^1\text{n}_0 +$  \_\_\_\_\_

# Key Terms Nuclear Changes Worksheet

**Yicheng Fang**

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## **Table of Contents Key Terms Nuclear Changes Worksheet**

1. Understanding the eBook Key Terms Nuclear Changes Worksheet
  - The Rise of Digital Reading Key Terms Nuclear Changes Worksheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Key Terms Nuclear Changes Worksheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Key Terms Nuclear Changes Worksheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Key Terms Nuclear Changes Worksheet
  - Personalized Recommendations
  - Key Terms Nuclear Changes Worksheet User Reviews and Ratings
  - Key Terms Nuclear Changes Worksheet and Bestseller Lists
5. Accessing Key Terms Nuclear Changes Worksheet Free and Paid eBooks
  - Key Terms Nuclear Changes Worksheet Public Domain eBooks
  - Key Terms Nuclear Changes Worksheet eBook Subscription Services
  - Key Terms Nuclear Changes Worksheet Budget-Friendly Options
6. Navigating Key Terms Nuclear Changes Worksheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Key Terms Nuclear Changes Worksheet Compatibility with Devices
  - Key Terms Nuclear Changes Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Key Terms Nuclear Changes Worksheet
  - Highlighting and Note-Taking Key Terms Nuclear Changes Worksheet
  - Interactive Elements Key Terms Nuclear Changes Worksheet

8. Staying Engaged with Key Terms Nuclear Changes Worksheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Key Terms Nuclear Changes Worksheet
9. Balancing eBooks and Physical Books Key Terms Nuclear Changes Worksheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Key Terms Nuclear Changes Worksheet
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Key Terms Nuclear Changes Worksheet
  - Setting Reading Goals Key Terms Nuclear Changes Worksheet
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Key Terms Nuclear Changes Worksheet
  - Fact-Checking eBook Content of Key Terms Nuclear Changes Worksheet
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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