

UNIT 1

11.1

- Condensed phases- a phase in which the particles are packed in close proximity to one another
- IMFs: solid>liquid>gas
- IMFs are the forces that exist between the molecules, atoms, or ions that make up a particular substance
- Stronger the IMFS, that harder it is to separate the particles.
- Fusion: solid $\rightarrow$ liquid and Vaporization: liquid $\rightarrow$ gas both require an input of energy
- Condensation: gas $\rightarrow$ liquid releases energy
- Enthalpy of vaporization ΔH_{vap} is the energy required to convert one mole of a liquid to vapor at a given temperature
- As IMF $\uparrow = \Delta H_{\text{vap}} \uparrow$



11.2

- Vapor Pressure- the pressure of the vapor of a substance in contact with its liquid or solid phase in a sealed container
 - When liquid is placed in a sealed container, some molecules escape into vapor phase (increasing vapor pressure) and some will recondense into liquid phase. When these become equal, they reach the dynamic equilibrium
 - Equilibrium vapor pressure is the pressure of vapor of a substance at equilibrium in contact with its liquid or solid phase in a sealed container
- $\uparrow$ temperature = $\uparrow$ VP
- $\uparrow$ IMFs = $\downarrow$ VP
- Boiling Point of a liquid is when the Vapor pressure of the liquid is equal to the external pressure
 - Normal boiling point is the temperature when the vapor pressure of a liquid is equal to 1 atm (760 mmHg)
- $\uparrow$ IMFs = $\uparrow$ BP

$$\ln P = \frac{-\Delta H_{\text{vap}}}{RT} + C \quad \ln \frac{P_2}{P_1} = \frac{-\Delta H_{\text{vap}}}{R} \left(\frac{1}{T_2} - \frac{1}{T_1} \right)$$

- $R = 8.314 \times 10^{-3} \text{ kJ/K} \cdot \text{mol}$, $C = \text{constant of integration}$
 - $\ln P$ vs. $1/T$ and the slope is $-\Delta H_{\text{vap}}/R$

11.3

- Cohesive forces: a force of attraction between molecules of a single substance and affect the physical properties of a liquid
- Surface tension is the measure of force required to "break" the surface of a liquid
 - Results from different IMFs experienced by molecules at the surface of a liquid and those in the interior of the liquid

General Chemistry Semester 2 Study Guide

Didier Musso



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intellectual growth Next she presents a specific study system that can quickly empower students to maximize their learning Then she addresses the importance of dealing with emotion attitudes and motivation by suggesting ways to change students mindsets about ability and by providing a range of strategies to boost motivation and learning finally she offers guidance to faculty on partnering with campus learning centers She pays particular attention to academically unprepared students noting that the strategies she offers for this particular population are equally beneficial for all students While stressing that there are many ways to teach effectively and that readers can be flexible in picking and choosing among the strategies she presents Saundra McGuire offers the reader a step by step process for delivering the key messages of the book to students in as little as 50 minutes Free online supplements provide three slide sets and a sample video lecture This book is written primarily for faculty but will be equally useful for TAs tutors and learning center professionals For readers with no background in education or cognitive psychology the book avoids jargon and esoteric theory

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