

General Chemistry

1st Semester Final Review

The final is 100 multiple choice questions over chapters 1-5. Approximately 25% of the test is mathematical calculations. There will be between 15-25 questions on each of the chapters

Chapter 1- Matter and Change

1. Know how to identify an intensive physical, extensive physical, and chemical property. Differentiate between physical and chemical reactions.
 - a. List two chemical and four physical properties of Iron
 - b. Identify the following as physical or chemical changes
 1. Baking a cake, melting candle wax, dissolving salt in water, garbage rotting, corrosion of metals, crushing a crystal to powder
 - c. Classify these properties: conductivity, ductility, corrosiveness, mass, density, combustibility. (Chemical Intensive, or Extensive physical)
2. Know the differences between elements, compounds, and the two types of mixtures (Homogeneous & Heterogeneous)
 - a. Compare and contrast mixtures and pure substances
 - b. Give two examples of homogeneous and heterogeneous mixtures
3. Know the difference between Exothermic and Endothermic and identify which type is occurring in a given reaction
 - a. A combustion Reaction like wood burning.
 - b. A Medical Ice Pack is snapped and it gets cold.
 - c. Ice Melts
4. Know the basic parts on the Periodic Table
 - a. Identify where metals, nonmetals, and metalloids are found
 - b. A neutral atom with 46 protons is a part of what group?
 - c. Identify the Families and what row they are in (include their ion charge)
 - d. An element in the Alkaline Earth Metals will take what charge?
 - e. What must nitrogen do to become stable? What charge will that make it?

Chapter 2- Measurement and Calculations

1. Know the difference between Accuracy and Precision
 - a. A set of students measure a test tube's mass to be 10.4g, 10.0g, 10.34g, 9.5g is the data as a whole precise if the actual mass is 8.0 grams?
 - b. Using the information above, tell which student was the most accurate. Which one was the most precise?
2. Know how to use significant digits
 - a. Calculate the following and round to proper significant digits (round answers to proper sigs)
 1. $(3.40 \times 10^{-5}) (.00285) / 127.66 =$
 2. $100.0 + 32.98 + 4.2 =$
 3. $(6.626 \times 10^{-34}) (2.1 \times 10^{26}) =$
3. Know how to use dimensional analysis to calculate
 - a. 320 nm to cm
 - b. 4.0×10^{-5} Mhz to Htz

General Chemistry 1st Semester Final Exam

Albert A Gayle



General Chemistry 1st Semester Final Exam:

Whispering the Techniques of Language: An Psychological Quest through **General Chemistry 1st Semester Final Exam**

In a digitally-driven earth wherever displays reign great and quick transmission drowns out the subtleties of language, the profound techniques and psychological subtleties concealed within words usually move unheard. Yet, located within the pages of **General Chemistry 1st Semester Final Exam** a interesting literary value sporting with fresh thoughts, lies a fantastic journey waiting to be undertaken. Published by a skilled wordsmith, that marvelous opus encourages visitors on an introspective trip, softly unraveling the veiled truths and profound impact resonating within the very fabric of each word. Within the mental depths of this touching review, we can embark upon a honest exploration of the book is primary styles, dissect its fascinating writing design, and yield to the strong resonance it evokes strong within the recesses of readers hearts.

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