

CHEM 121: Final Exam Study Guide

Chapter 1

- Know the Scientific method
- Know the definitions for hypothesis, scientific law, and scientific theory.

Chapter 2

- Length, mass, weight, volume
 - Know $1\text{ cm}^3=1\text{ mL}$ and $1\text{ dm}^3=1\text{ L}$
- Significant figures and Scientific notation
 - Rounding, in addition or subtraction, in multiplication or division
 - Keep as many sig figs until final answer
- Scientific notation
- Metric system
 - Know the metric prefixes: kilo (k), deci (d), centi (c), milli (m), and micro (μ)
 - Be able to perform metric-metric conversions using these prefixes.
- Use the metric-English conversions provided ($1\text{ in.}=2.54\text{ cm}$; $1\text{ lb}=453.6\text{ g}$; $1\text{ qt}=946\text{ mL}$)
- Volume by calculation
 - $V_{\text{rectangular solid}} = \text{length} \times \text{width} \times \text{thickness}$
- Volume by displacement
- Density: $d = \frac{m}{V}$
 - Calculate density, mass, or volume
 - Identify what items sink or float given densities of liquids and solids.
- Temperature
 - Know the formulas for converting "F-to-"C or "C-to-"F and K-to-"C or "C-to-K
- Percentage: ratio of parts per 100 parts
 - Given amount of part and whole, calculate %
 - Use a given % to solve for part or whole

Chapter 3

- Know that matter is studied at the macroscopic, microscopic, particulate (molecular) levels
- Physical states of matter
 - Determine physical state of substances (solids, liquids, gases) given descriptions of volume, shape, particles moving, etc.
- Classify properties and changes as physical or chemical
 - Know terms for changes of state:
 - Melting, freezing, vaporizing, condensation, sublimation, deposition
- Classification of matter
 - Given examples, determine which are elements, compounds, or mixtures
 - Given molecular-level images, determine which are elements, compounds, or mixtures and solids, liquids, or gases
 - Distinguish between homogeneous and heterogeneous mixtures
- Chemical reaction:
 - reactants: starting materials
 - products: substances produced in reaction

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Table of Contents Final Exam Study Guide Chemistry

1. Understanding the eBook Final Exam Study Guide Chemistry
 - The Rise of Digital Reading Final Exam Study Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Final Exam Study Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Final Exam Study Guide Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Final Exam Study Guide Chemistry
 - Personalized Recommendations

- Final Exam Study Guide Chemistry User Reviews and Ratings
- Final Exam Study Guide Chemistry and Bestseller Lists
- 5. Accessing Final Exam Study Guide Chemistry Free and Paid eBooks
 - Final Exam Study Guide Chemistry Public Domain eBooks
 - Final Exam Study Guide Chemistry eBook Subscription Services
 - Final Exam Study Guide Chemistry Budget-Friendly Options
- 6. Navigating Final Exam Study Guide Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Final Exam Study Guide Chemistry Compatibility with Devices
 - Final Exam Study Guide Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Final Exam Study Guide Chemistry
 - Highlighting and Note-Taking Final Exam Study Guide Chemistry
 - Interactive Elements Final Exam Study Guide Chemistry
- 8. Staying Engaged with Final Exam Study Guide Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Final Exam Study Guide Chemistry
- 9. Balancing eBooks and Physical Books Final Exam Study Guide Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Final Exam Study Guide Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Final Exam Study Guide Chemistry
 - Setting Reading Goals Final Exam Study Guide Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Final Exam Study Guide Chemistry
 - Fact-Checking eBook Content of Final Exam Study Guide Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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