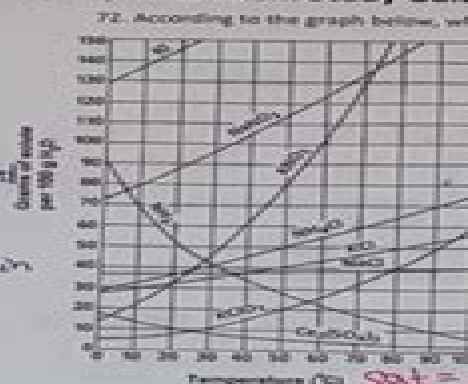


### Vocab

64. Unsaturated
65. Saturated  
maximum amt. of solute dissolved  
at that temp.
66. Solubility  
the amount (g) of solute  
that can dissolve at a certain  
temp.
67. Rate of dissolution  
how fast a solute can  
dissolve
68. Molarity  
concentration  $M = \frac{\text{mol solute}}{1 \text{ Liter}}$
69. Supersaturated  
more than max amt. of solute  
dissolved in solvent
70. Electrolyte  
- must be soluble  
- must be ionic  
can conduct electricity
71. List 3 ways to increase the rate of dissolution of a solid:
1. increase surface area (crush)
  2. increase temperature
  3. stir

### SOLUBILITY OF COMMON IONIC COMPOUNDS IN WATER

<u><b>Soluble compounds contain</b></u>	<u><b>Common exceptions</b></u>
$C_2H_3O_2^-$ , $CH_3COO^-$	None
$NH_4^+$	None
$NO_3^-$	None
$Cl^-$	None
$ClO_3^-$	None
$ClO_4^-$	None
$Br^-$	Compounds of $Ag^+$ , $Pb^{2+}$ , and $Hg_2^{2+}$
$I^-$	Compounds of $Ag^+$ , $Pb^{2+}$ , and $Hg_2^{2+}$
$SO_4^{2-}$	Compounds of $Ag^+$ , $Pb^{2+}$ , and $Hg_2^{2+}$
$SO_3^{2-}$	Compounds of $Sr^{2+}$ , $Ba^{2+}$ , $Pb^{2+}$ , and $Hg_2^{2+}$
<u><b>Insoluble compounds contain</b></u>	<u><b>Common exceptions</b></u>
$CO_3^{2-}$	Compounds of $Na^+$ and the alkali metal cations
$PO_4^{3-}$	Compounds of $Na^+$ and the alkali metal cations
$CrO_4^{2-}$	Compounds of $Na^+$ and the alkali metal cations
$Cr_2O_7^{2-}$	Compounds of $Na^+$ and the alkali metal cations
$OH^-$	Compounds of $Na^+$ , the alkali metal cations, $Ca^{2+}$ , $Sr^{2+}$ , and $Ba^{2+}$
$S^{2-}$	Compounds of $Na^+$ , the alkali metal cations, $Ca^{2+}$ , $Sr^{2+}$ , and $Ba^{2+}$



76. Use the solubility rules table on the left to help you with the following:

	Soluble	Insoluble	Dissociation reaction	Is it an electrolyte?	How many ions are produced?
$\text{NaNO}_3$	✓		$\text{NaNO}_3 \xrightarrow{\text{aq}} \text{Na}^+ + \text{NO}_3^-$	✓	2
$\text{Na}_2\text{CO}_3$	✓		$\text{Na}_2\text{CO}_3 \xrightarrow{\text{aq}} 2\text{Na}^+ + \text{CO}_3^{2-}$	✓	3
$\text{NaCl}$	✓		$\text{NaCl} \xrightarrow{\text{aq}} \text{Na}^+ + \text{Cl}^-$	✓	2
$\text{Ca}(\text{NO}_3)_2$	✓		$\text{Ca}(\text{NO}_3)_2 \xrightarrow{\text{aq}} \text{Ca}^{2+} + 2\text{NO}_3^-$	✓	3
$\text{CaCO}_3$		✓	does not dissociate	no	0

Molarity (M) is the number of moles of solute in one liter of solution

$$\text{molarity}(M) = \frac{\text{amount of solute (mol)}}{\text{volume of solution (L)}}$$

**Dilution:** the action of making a stock solution less concentrated

$$M_{11}V_1 = M_{21}V_2$$

77. Calculate the Molarity of a solution that contains 100grams of NaCl in 5.00L of solution( the molar mass of NaCl is 58.44 g/mol)

① convert g  $\rightarrow$  mol  
 $\frac{100 \text{ g NaCl}}{58.44 \text{ g}} \times \frac{1 \text{ mol}}{1} = 1.71 \text{ mol}$

$$m = \frac{\text{mol}}{L}$$

$$M = \frac{1.71 \text{ mol}}{5.00 \text{ L}} =$$

0.34 molar solution NaCl

# Final Exam Review Packet Answers Chemistry

**Jay Wackerly, Sarah Zingales, Michael  
Wentzel, Gautam Bhattacharyya, Brett  
McCollum**

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practitioners and 3 highlight how people have used chemical education based research in their teaching practice There are places in the literature dedicated for chemical education research CER however there is not a clear avenue for those that have changed their teaching methods based on published CER and report their experiences Creating this article collection will foster collaboration between chemical education researchers and teachers of organic chemistry This opportunity allows these instructors to share evidence based practices experiences challenges and innovative approaches from CER literature and beyond This Research Topic bridges discipline based education research and the scholarship of teaching and learning which will help advance organic chemistry education and improve student outcomes **Phlebotomy Exam Review** Ruth E. McCall, Cathee M. Tankersley, 2004 This exam review now in its Second Edition prepares students for the major national certification exams in phlebotomy including ASCP NCA ASPT NPA and AMT In a comprehensive outline format it includes content review and contains over 800 multiple choice questions with answers and rationale Also includes sample exams to encourage a comfort level in the test taking environment A CD ROM with additional 150 sample test questions accompanies this helpful review manual

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