

Substance	Lewis Structure	Molar Mass	Intermolecular Attractive Force
Methanol CH_3OH $12 + 3(1) + 16 + 1 = 32$	$\begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \ddot{\text{O}} - \text{H} \\ \\ \text{H} \end{array}$	32.042	Hydrogen Bonds
Ethanol $\text{CH}_3\text{CH}_2\text{OH}$ $12 + 5(1) + 16 + 1 = 46$	$\begin{array}{c} \text{H} \quad \text{H} \quad \\ \quad \quad \\ \text{H} - \text{C} - \text{C} - \ddot{\text{O}} - \text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$	46.068	Hydrogen Bonds
n-Butanol $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$	$\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \quad \quad \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{C} - \ddot{\text{O}} - \text{H} \\ \quad \quad \quad \\ \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \end{array}$	74.122	Hydrogen Bonds
Glycerin $\text{C}_3\text{H}_8(\text{OH})_3$ $12 + 5(1) + 18(16) = 92$	$\begin{array}{c} \text{H} \quad \text{:}\ddot{\text{O}}\text{:} - \text{H} \quad \text{H} \\ \quad \quad \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{H} \\ \quad \quad \\ \text{:}\ddot{\text{O}}\text{:} - \text{H} \quad \text{H} \quad \text{:}\ddot{\text{O}}\text{:} - \text{H} \end{array}$	92.093	Hydrogen Bonds
Water H_2O	$\begin{array}{c} \text{H} \\ \\ \text{H} - \ddot{\text{O}} - \text{H} \\ \\ \text{H} \end{array}$	18.015	Hydrogen Bonds

Exploring Intermolecular Forces Lab Answers

Klaus D. Sattler



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Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

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Exploring Chemistry (Loose-Leaf) Matthew Johll, 2012-05-04 Matthew Johll's *Exploring Chemistry* covers the standard topics for the nonmajors course in the typical order but each chapter unfolds in the context of a single case study that helps students connect what they are learning to real life situations For example students work through the often difficult topics of molecular structure gas laws and organic chemistry by learning about the development of powerful new chemotherapy drugs new technologies for screening airline passengers and the creation of biodegradable biopolymers It's the same case driven approach that Johll uses in his acclaimed *Investigating Chemistry* now in its Third Edition but *Exploring Chemistry* goes beyond the other book's specific focus on examples from forensic science to use real life stories from cooking athletics genetics green chemistry and more

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