

Electron Configurations

Reference Page

Electron Configuration: A way to show the arrangement of electrons in an atom.

Periodic Table: A grouping of elements based on their atomic number.

Periodic table groups elements together based on their atomic number. Elements in the same group have similar properties because they have the same number of valence electrons.

1. **Electron shells:** Electrons are found in shells around the nucleus. Shells are arranged from the innermost shell to the outermost shell. The innermost shell can hold up to 2 electrons. The next shell can hold up to 8 electrons. The outermost shell can hold up to 8 electrons.
2. **Electron subshells:** Electrons are found in subshells within each shell. Subshells are arranged from the innermost shell to the outermost shell. The innermost shell can hold up to 2 electrons. The next shell can hold up to 8 electrons. The outermost shell can hold up to 8 electrons.
3. **Electron configurations:** An arrangement showing the arrangement of the atoms electrons. Electrons follow the Aufbau principle which states that electrons will fill the lowest energy level first.



Electron shells

Each shell represents another ring of electrons orbiting the nucleus. Shells are numbered starting from the innermost shell.

The shells are numbered 1 through 7, starting from the innermost shell.

1. **Shell 1:** Can hold up to 2 electrons. This shell is also called the K shell.
2. **Shell 2:** Can hold up to 8 electrons. This shell is also called the L shell.
3. **Shell 3:** Can hold up to 8 electrons. This shell is also called the M shell.
4. **Shell 4:** Can hold up to 8 electrons. This shell is also called the N shell.
5. **Shell 5:** Can hold up to 8 electrons. This shell is also called the O shell.
6. **Shell 6:** Can hold up to 8 electrons. This shell is also called the P shell.
7. **Shell 7:** Can hold up to 8 electrons. This shell is also called the Q shell.



Electrons fill the shells in order of increasing atomic number. Electrons fill the shells in order of increasing atomic number.

1. **Filling the shells:** Electrons fill the shells in order of increasing atomic number.
2. **Outer shell:** An atom's outer shell contains the valence electrons.



Explore Learning Electron Configuration Answer Key

Michael Brown

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