

### Limiting Reactant Practice 2.3 - Answer Key

**Objective:** Use stoichiometry to identify the limiting and excess reactants.

**Directions:** Answer the following limiting reactant problems. Show your work!!

1. A reaction between potassium and oxygen gas produces potassium oxide, represented by the following balanced chemical equation:



- a. If 3.00 g of potassium and 5.00 g of oxygen are mixed, how much potassium oxide is produced?

$$3.00 \text{ g K} \times \left( \frac{1 \text{ mol K}}{39.10 \text{ g K}} \right) \left( \frac{1 \text{ mol O}_2}{2 \text{ mol K}} \right) \left( \frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = \boxed{1.21 \text{ g K}_2\text{O}}$$

$$5.00 \text{ g O}_2 \times \left( \frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right) \left( \frac{1 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) \left( \frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = 5.00 \text{ g K}_2\text{O}$$

- b. What is the limiting reactant?

Potassium (K)

- c. What is the reactant in excess?

Oxygen (O<sub>2</sub>)

2. In a reaction, 12.00 g of aqueous sodium hydroxide reacts with aqueous 22.00 g of copper(II) sulfate.



- a. What mass of sodium sulfate is produced?

$$12.00 \text{ g NaOH} \times \left( \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \right) \left( \frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol NaOH}} \right) \left( \frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 17.10 \text{ g}$$

$$22.00 \text{ g CuSO}_4 \times \left( \frac{1 \text{ mol CuSO}_4}{159.61 \text{ g CuSO}_4} \right) \left( \frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) \left( \frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = \boxed{19.87 \text{ g}}$$

- b. What is the limiting reactant?

Copper(II) sulfate (CuSO<sub>4</sub>)

- c. What is the reactant in excess?

Sodium Hydroxide (NaOH)

# Limiting Reactants Study Guide Answers

**J Elliott**



## Limiting Reactants Study Guide Answers:

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Dolter,2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

**CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE** Kevin Smith,2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text *Chemistry, Student Study Guide* James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

**Chemical Principles Study Guide/Solutions Manual** John Krenos,Joseph Potenza,Carl Hoeger,2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

**Study Guide to Accompany Calculus for the Management, Life, and Social Sciences** Clyde Metz,1984-01-01 Study Guide to Accompany Calculus for the Management Life and Social Sciences *Chemical Principles Student's Study Guide & Solutions Manual* John Krenos,Joseph Potenza,Loretta Jones,Lynn Koplitiz,Thomas

Spence,2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the

material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

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