

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

Limiting Reactant And Limiting Yield Answer Key

**John A. Olmsted, Gregory M.
Williams, Robert Charles Burk**



Limiting Reactant And Limiting Yield Answer Key:

The book delves into Limiting Reactant And Limiting Yield Answer Key. Limiting Reactant And Limiting Yield Answer Key is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Limiting Reactant And Limiting Yield Answer Key, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Limiting Reactant And Limiting Yield Answer Key
 - Chapter 2: Essential Elements of Limiting Reactant And Limiting Yield Answer Key
 - Chapter 3: Limiting Reactant And Limiting Yield Answer Key in Everyday Life
 - Chapter 4: Limiting Reactant And Limiting Yield Answer Key in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, the author will provide an overview of Limiting Reactant And Limiting Yield Answer Key. This chapter will explore what Limiting Reactant And Limiting Yield Answer Key is, why Limiting Reactant And Limiting Yield Answer Key is vital, and how to effectively learn about Limiting Reactant And Limiting Yield Answer Key.
 3. In chapter 2, the author will delve into the foundational concepts of Limiting Reactant And Limiting Yield Answer Key. The second chapter will elucidate the essential principles that need to be understood to grasp Limiting Reactant And Limiting Yield Answer Key in its entirety.
 4. In chapter 3, the author will examine the practical applications of Limiting Reactant And Limiting Yield Answer Key in daily life. The third chapter will showcase real-world examples of how Limiting Reactant And Limiting Yield Answer Key can be effectively utilized in everyday scenarios.
 5. In chapter 4, the author will scrutinize the relevance of Limiting Reactant And Limiting Yield Answer Key in specific contexts. This chapter will explore how Limiting Reactant And Limiting Yield Answer Key is applied in specialized fields, such as education, business, and technology.
 6. In chapter 5, this book will draw a conclusion about Limiting Reactant And Limiting Yield Answer Key. This chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Limiting Reactant And Limiting Yield Answer Key.

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