






### Stoichiometry Question

Consider the synthesis of water produced in the Model. The reaction container is filled with 10.0 g of  $H_2$  and 1.0 g of  $O_2$ .

- a. Which reactant is the degree of progress in the limiting reactant in this case? Show your work. Also indicate that you are given standard quantities in mass units (g) – not moles.

$$10.0 \text{ g } H_2 \left( \frac{1 \text{ mol } H_2}{2.016 \text{ g}} \right) = 4.96 \text{ mol } H_2$$

$$1.0 \text{ g } O_2 \left( \frac{1 \text{ mol } O_2}{32.00 \text{ g}} \right) = 0.031 \text{ mol } O_2$$

	$2 H_2 +$	$O_2 \rightarrow$	$2 H_2O$
Before	10.0 g $H_2$ 	1.0 g $O_2$ 	0.00 g $H_2O$
Change 10.0 g $H_2$ 1.0 g $O_2$ 0.00 g $H_2O$ 0.00 g $H_2O$	-10.0 g $H_2$ -1.0 g $O_2$ 0.00 g $H_2O$ 0.00 g $H_2O$	+1.0 g $O_2$ +1.0 g $O_2$ 0.00 g $H_2O$ 0.00 g $H_2O$	-10.0 g $H_2$ -1.0 g $O_2$ 0.00 g $H_2O$ 0.00 g $H_2O$
After	0.00 g $H_2$ 	0.00 g $O_2$ 	0.00 g $H_2O$ 

Change in the limiting reactant



- b. What mass of water can be produced? Show your work.

$$0.031 \text{ mol } O_2 \left( \frac{2 \text{ mol } H_2O}{1 \text{ mol } O_2} \right) = 0.062 \text{ mol } H_2O$$

- c. Which reactant is present in excess, and what mass of that reactant remains after the reaction is complete?

$H_2$  is present in excess  
initial mass of  $H_2$  remains

$$10.0 \text{ g } H_2 \left( \frac{2.016 \text{ g}}{1 \text{ mol}} \right) = 4.96 \text{ g } H_2$$

# Limiting And Excess Reactants Extension Questions Pogil

**L Cohen**



## **Limiting And Excess Reactants Extension Questions Pogil:**

## Decoding **Limiting And Excess Reactants Extension Questions Pogil**: Revealing the Captivating Potential of Verbal Expression

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