

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PerV=nRT$ ", and the universal gas constant $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 kPa / 1 atm$ to get
 $R = 8.31 kPa \cdot L / (K \cdot mole)$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 atm)(12 L)}{4 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 K$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of $87^{\circ}C$, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 atm \cdot 31 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}$$

$$n = 1.2586 mol$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 K}{60 L}$$

$$P = 1.642 atm$$

$$or \quad P = 166.29 kPa$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of $56^{\circ}C$, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 327 K}{0.09 atm}$$

$$V = 2310.93 L$$

- 5) If I have 17 moles of gas at a temperature of $67^{\circ}C$, and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 390 K}{88.89 L}$$

$$P = 5.34 atm$$

$$or \quad P = 540.61 kPa$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 atm \cdot 25 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 K}$$

$$n = 0.5075 mol$$

Gases Properties And Laws Answers

Jin-Ying Zhang



Gases Properties And Laws Answers:

Engineering Thermodynamics Dr. S. S. Khandare, 2003 This book on Engineering Thermodynamic contains basic principles and fundamental laws of Thermal Engineering It deals with the gas laws and properties of fluids like pressure temperature and volume The book discusses the thermodynamic processes like isothermal isentropic and polytropic processes The new concept of availability and irreversibility has been included in the book The various properties like enthalpy entropy internal energy of steam are discussed The topics on properties of steam and steam cycles like rankine modified rankine cycles are also presented in the book *Physical Chemistry for JEE Advanced: Part 1, 3E (Free Sample)* K. S. Verma, 2022-05-19 Physical Chemistry for JEE Advanced Part 1 a Cengage Exam Crack Series product is designed to help aspiring engineers focus on the subject of physical chemistry from two standpoints To develop their caliber aptitude and attitude for the engineering field and profession To strengthen their grasp and understanding of the concepts of the subjects of study and their applicability at the grassroots level Each book in this series approaches the subject in a very conceptual and coherent manner While its illustrative solved examples facilitate easy mastering of the concepts and their applications an array of solved problems exposes the students to a variety of questions that they can expect in the examination The coverage and features of this series of books make it highly useful for all those preparing for JEE Main and Advanced and aspiring to become engineers *Oil & Gas Journal*, 1928 **Utility Corporations** United States. Federal Trade Commission, 1936

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Table of Contents Gases Properties And Laws Answers

1. Understanding the eBook Gases Properties And Laws Answers
 - The Rise of Digital Reading Gases Properties And Laws Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gases Properties And Laws Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gases Properties And Laws Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gases Properties And Laws Answers

- Personalized Recommendations
- Gases Properties And Laws Answers User Reviews and Ratings
- Gases Properties And Laws Answers and Bestseller Lists
- 5. Accessing Gases Properties And Laws Answers Free and Paid eBooks
 - Gases Properties And Laws Answers Public Domain eBooks
 - Gases Properties And Laws Answers eBook Subscription Services
 - Gases Properties And Laws Answers Budget-Friendly Options
- 6. Navigating Gases Properties And Laws Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gases Properties And Laws Answers Compatibility with Devices
 - Gases Properties And Laws Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gases Properties And Laws Answers
 - Highlighting and Note-Taking Gases Properties And Laws Answers
 - Interactive Elements Gases Properties And Laws Answers
- 8. Staying Engaged with Gases Properties And Laws Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gases Properties And Laws Answers
- 9. Balancing eBooks and Physical Books Gases Properties And Laws Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gases Properties And Laws Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gases Properties And Laws Answers
 - Setting Reading Goals Gases Properties And Laws Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gases Properties And Laws Answers

- Fact-Checking eBook Content of Gases Properties And Laws Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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