

Gas Laws Worksheet

$$\text{atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

Boyle's Law Problems:

$$P_1 V_1 = P_2 V_2$$

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(748 \text{ mmHg})(22.5 \text{ L}) = (725 \text{ mmHg}) V_2$$

$$V_2 = \frac{(748 \text{ mmHg})(22.5 \text{ L})}{(725 \text{ mmHg})}$$

$$V_2 = 23.2 \text{ L}$$

2. A gas with a volume of 4.0 L at a pressure of 205 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?

$$(4.0 \text{ L})(205 \text{ kPa}) = (12.0 \text{ L}) P_2$$

$$P_2 = \frac{(4.0 \text{ L})(205 \text{ kPa})}{12.0 \text{ L}}$$

$$P_2 = 68.3 \text{ kPa}$$

3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

$$(196.0 \text{ L})(1.00 \text{ atm}) = (26.0 \text{ L}) P_2$$

$$P_2 = \frac{(196.0 \text{ L})(1.00 \text{ atm})}{26.0 \text{ L}}$$

$$P_2 = 7.54 \text{ atm}$$

4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

$$(40.0 \text{ L})(12.7 \text{ kPa}) = (8.4 \text{ kPa}) V_2$$

$$V_2 = \frac{(40.0 \text{ L})(12.7 \text{ kPa})}{8.4 \text{ kPa}}$$

$$V_2 = 60.5 \text{ L}$$

Gas Laws Practice Calculations Unit 9 Lesson 1

Ying-Ying Zheng

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Gas Laws Practice Calculations Unit 9 Lesson 1:

Unveiling the Energy of Verbal Art: An Mental Sojourn through **Gas Laws Practice Calculations Unit 9 Lesson 1**

In a world inundated with screens and the cacophony of instantaneous communication, the profound energy and emotional resonance of verbal artistry frequently diminish in to obscurity, eclipsed by the constant assault of sound and distractions. Yet, located within the lyrical pages of **Gas Laws Practice Calculations Unit 9 Lesson 1**, a fascinating perform of literary splendor that impulses with natural emotions, lies an remarkable trip waiting to be embarked upon. Published with a virtuoso wordsmith, that interesting opus instructions viewers on a psychological odyssey, softly exposing the latent possible and profound influence embedded within the complex internet of language. Within the heart-wrenching expanse with this evocative evaluation, we can embark upon an introspective exploration of the book is main subjects, dissect its captivating writing model, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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