

Mixed Gas Law Worksheet

Answers available at end

Name: _____ (show your work) period: _____

Convert:

$$100 \text{ mmHg} = \text{_____ kPa}$$

$$300 \text{ mmHg} = \text{_____ atm}$$

$$120 \text{ atm} = \text{_____ mmHg}$$

$$300 \text{ torr} = \text{_____ kPa}$$

$$25 \text{ kPa} = \text{_____ mmHg}$$

$$25 \text{ mmHg} = \text{_____ atm}$$

$$300^{\circ}\text{C} = \text{_____ K}$$

$$150\text{K} = \text{_____}^{\circ}\text{C}$$

$$5000^{\circ}\text{C} = \text{_____ K}$$

$$300\text{K} = \text{_____}^{\circ}\text{C}$$

$$-255^{\circ}\text{C} = \text{_____ K}$$

$$3500\text{K} = \text{_____}^{\circ}\text{C}$$

$$200^{\circ}\text{C} = \text{_____}^{\circ}\text{F}$$

$$95^{\circ}\text{F} = \text{_____}^{\circ}\text{C} = \text{_____ K}$$

$$45^{\circ}\text{F} = \text{_____}^{\circ}\text{C} = \text{_____ K}$$

- 1) A 70. liter sample of gas initially at 280°C is allowed to cool at constant pressure, what will the new volume be at 0°C ?

$$P_1 = \text{_____} \quad P_2 = \text{_____}$$

$$V_1 = \text{_____} \quad V_2 = \text{_____}$$

$$T_1 = \text{_____} \quad T_2 = \text{_____}$$

- 2) A 4.0 liter sample of gas initially at 25°C is allowed to cool at constant pressure, what will the new volume be at -5°C ?

$$P_1 = \text{_____} \quad P_2 = \text{_____}$$

$$V_1 = \text{_____} \quad V_2 = \text{_____}$$

$$T_1 = \text{_____} \quad T_2 = \text{_____}$$

- 3) A buoyancy vest (used in scuba diving) is filled to a volume of 2.3 liters at a pressure of 1 atmosphere (at the surface where the temperature is a balmy 300K), what will the volume of air be in the vest when the diver dives to a depth of about 90 feet where the temperature is 17°C and the pressure is 4 atmospheres?

$$P_1 = \text{_____} \quad P_2 = \text{_____}$$

Gas Laws And Scuba Diving Worksheet Answers

Puja Mehta



Gas Laws And Scuba Diving Worksheet Answers:

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Table of Contents Gas Laws And Scuba Diving Worksheet Answers

1. Understanding the eBook Gas Laws And Scuba Diving Worksheet Answers
 - The Rise of Digital Reading Gas Laws And Scuba Diving Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws And Scuba Diving Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws And Scuba Diving Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws And Scuba Diving Worksheet Answers
 - Personalized Recommendations
 - Gas Laws And Scuba Diving Worksheet Answers User Reviews and Ratings

- Gas Laws And Scuba Diving Worksheet Answers and Bestseller Lists
- 5. Accessing Gas Laws And Scuba Diving Worksheet Answers Free and Paid eBooks
 - Gas Laws And Scuba Diving Worksheet Answers Public Domain eBooks
 - Gas Laws And Scuba Diving Worksheet Answers eBook Subscription Services
 - Gas Laws And Scuba Diving Worksheet Answers Budget-Friendly Options
- 6. Navigating Gas Laws And Scuba Diving Worksheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Laws And Scuba Diving Worksheet Answers Compatibility with Devices
 - Gas Laws And Scuba Diving Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws And Scuba Diving Worksheet Answers
 - Highlighting and Note-Taking Gas Laws And Scuba Diving Worksheet Answers
 - Interactive Elements Gas Laws And Scuba Diving Worksheet Answers
- 8. Staying Engaged with Gas Laws And Scuba Diving Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws And Scuba Diving Worksheet Answers
- 9. Balancing eBooks and Physical Books Gas Laws And Scuba Diving Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws And Scuba Diving Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Laws And Scuba Diving Worksheet Answers
 - Setting Reading Goals Gas Laws And Scuba Diving Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Laws And Scuba Diving Worksheet Answers
 - Fact-Checking eBook Content of Gas Laws And Scuba Diving Worksheet Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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