

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Instructions

Standard Conditions

Unit	Standard Pressure
atm	1.013
torr	760
mm	760
cm	760
in	760

22.4 L = 1 mol at STP

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Gas Law	Equation
Boyle's Law	$P_1V_1 = P_2V_2$
Charles's Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$
Gay-Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$
Avogadro's Law	$\frac{V_1}{n_1} = \frac{V_2}{n_2}$
Combined Gas Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$
Ideal Gas Law	$PV = nRT$

The constant R (used in the ideal gas law) depends on the unit of pressure:

- 0.0821 L·atm/(mol·K)
- 8.314 J/(mol·K)

Notes:

- Temperature must always be in Kelvin! $K = ^\circ C + 273$
- If you are given two different units of pressure, convert them.
- Create two columns before solving: initial and final. Write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (For Ideal Gas Law, write what you're given, insert those in equation and solve for unknown.)
- Avogadro's Law: DON'T FORGET TO USE MOLE! (Mol/L)
- Combined Gas Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effluent and molecular mass in the correct spots before solving the equation. (This is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

Charles' Law $V_1/T_1 = V_2/T_2$

1. The temperature under my refrigerator is about 4°C. If I place it below the temperature of 4°C, will it become colder, warmer, or stay the same? Explain.

2. A gas is at 100°C and 1 atm. If I change the temperature to 200°C, what happens to the volume? Explain.

Gay-Lussac's Law $P_1/T_1 = P_2/T_2$

1. A gas is at 100°C and 1 atm. If I change the temperature to 200°C, what happens to the pressure? Explain.

2. A gas is at 100°C and 1 atm. If I change the temperature to 200°C, what happens to the volume? Explain.

Avogadro's Law $V_1/n_1 = V_2/n_2$

1. A gas is at 100°C and 1 atm. If I change the temperature to 200°C, what happens to the volume? Explain.

2. A gas is at 100°C and 1 atm. If I change the temperature to 200°C, what happens to the volume? Explain.

Combined Gas Law $\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$

1. A gas is at 100°C and 1 atm. If I change the temperature to 200°C, what happens to the volume? Explain.

2. A gas is at 100°C and 1 atm. If I change the temperature to 200°C, what happens to the volume? Explain.

Ideal Gas Law $PV = nRT$

1. A gas is at 100°C and 1 atm. If I change the temperature to 200°C, what happens to the volume? Explain.

2. A gas is at 100°C and 1 atm. If I change the temperature to 200°C, what happens to the volume? Explain.

All Gas Laws Included!

Gas Law Problems Packet

Kalliat T. Valsaraj, Elizabeth M. Melvin

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the authors' names.

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