

Answers to Solubility Rules Worksheet

11. Classify each ionic compound as being soluble or insoluble in water.

- | | |
|--|---|
| <ul style="list-style-type: none"> a. potassium carbonate - SOLUBLE b. lead(II) carbonate - NOT SOLUBLE c. barium sulfate - NOT SOLUBLE d. zinc hydroxide - NOT SOLUBLE e. calcium carbonate - NOT SOLUBLE f. aluminum - NOT SOLUBLE g. ammonium chloride - NOT SOLUBLE h. zinc chloride - NOT SOLUBLE | <ul style="list-style-type: none"> i. silver acetate - SOLUBLE j. copper(II) sulfate - NOT SOLUBLE k. $\text{Fe}_3(\text{PO}_4)_2$ - NOT SOLUBLE l. NiSO_4 - SOLUBLE m. CuCl_2 - SOLUBLE n. CaSO_4 - NOT SOLUBLE o. Hg_2SO_4 - NOT SOLUBLE p. PbI_2 - NOT SOLUBLE |
|--|---|

12. Identify the reaction type(s) and precipitate formation, or no precipitate, for the following reactions, assuming complete ionic forms of the compounds that would give place to the products.

	KBr	Na_2CO_3	CuI_2	Na_2SO_4
AgNO_3	$\text{AgBr} + \text{KNO}_3$	$\text{NaNO}_3 + \text{Ag}_2\text{CO}_3$	$\text{Cu}(\text{NO}_3)_2 + \text{Ag}_2\text{S}$	$\text{AgOH} + \text{NH}_4\text{NO}_3$
BaCl_2	$\text{KCl} + \text{BaBr}_2$	$\text{NaCl} + \text{BaCO}_3$	$\text{CuCl}_2 + \text{BaS}$	$\text{Ba}(\text{OH})_2 + \text{NH}_4\text{Cl}$
$\text{Al}(\text{NO}_3)_3$	$\text{AlBr}_3 + \text{KNO}_3$	$\text{NaNO}_3 + \text{Al}_2(\text{CO}_3)_3$	$\text{Cu}(\text{NO}_3)_2 + \text{Al}_2\text{S}_3$	$\text{NH}_4\text{NO}_3 + \text{NH}_4\text{OH}$
CuSO_4	$\text{CuBr} + \text{KSO}_4$	$\text{Na}_2\text{SO}_4 + \text{CuCl}_2$	$\text{CuS} + \text{CuSO}_4$	$(\text{NH}_4)_2\text{SO}_4 + \text{Cu}(\text{OH})_2$

Experiment 26 Solubility Rules Answer Key

Rosina Ehmann



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