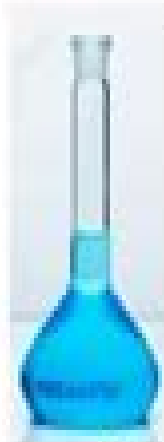




3.26g



80 cm³ / 0.2M



$$\begin{array}{c} 1 \quad \quad 1 \\ \quad \quad \times \\ 0.0583 \quad x \end{array}$$

$$(1)x = (0.0583)(1)$$

$$x = 0.0583 \text{ moles CuSO}_4$$

$$n = \frac{m}{A_r}$$

$$= \frac{3.26 \text{ g}}{55.85 \text{ g/mol}}$$

$$= 0.0583 \text{ moles}$$

$$C = \frac{n}{V(\text{dm}^3)}$$

$$0.2 = \frac{n}{0.08}$$

$$n = 0.2 \times 0.08 = 0.016 \text{ moles}$$

limiting reagent

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H Kauffman



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