

Limiting Reactant- Theoretical and Percentage Yields Key



If a reaction vessel contains 0.15 mol KO_2 and 0.10 mol H_2O , what is the limiting reactant? How many moles of oxygen can be produced?

$$0.15 \text{ mol KO}_2 \times \frac{3 \text{ mol O}_2}{4 \text{ mol KO}_2} = 0.11 \text{ mol O}_2$$

$$0.10 \text{ mol H}_2\text{O} \times \frac{3 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}} = 0.15 \text{ mol O}_2$$

Limiting reactant: KO_2

Maximum or theoretical yield = 0.11 mol O_2



If a reaction vessel contains 10.0 g of sodium chloride and 12.0 g of sulfuric acid, what is the limiting reactant? What is the theoretical yield of hydrochloric acid? How many grams of excess reactant are left unconsumed?



$$10.0 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} \times \frac{2 \text{ mol HCl}}{2 \text{ mol NaCl}} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 6.24 \text{ g HCl}$$

$$12.0 \text{ g H}_2\text{SO}_4 \times \frac{1 \text{ mol H}_2\text{SO}_4}{98.08 \text{ g H}_2\text{SO}_4} \times \frac{2 \text{ mol HCl}}{1 \text{ mol H}_2\text{SO}_4} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 8.92 \text{ g HCl}$$

Limiting reactant: NaCl

Maximum or theoretical yield = 6.24 g HCl

$$10.0 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} \times \frac{1 \text{ mol H}_2\text{SO}_4}{2 \text{ mol NaCl}} \times \frac{98.08 \text{ g H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} = 8.39 \text{ g H}_2\text{SO}_4 \text{ required to consume all of the NaCl (Limiting reactant)}$$

$12.0 \text{ g H}_2\text{SO}_4 - 8.39 \text{ g H}_2\text{SO}_4 = 3.61 \text{ g}$ of excess H_2SO_4 remains after reaction is complete

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Table of Contents Limiting Reagent Homework Answer Key

1. Understanding the eBook Limiting Reagent Homework Answer Key
 - The Rise of Digital Reading Limiting Reagent Homework Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent Homework Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagent Homework Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent Homework Answer Key
 - Personalized Recommendations

- Limiting Reagent Homework Answer Key User Reviews and Ratings
- Limiting Reagent Homework Answer Key and Bestseller Lists
- 5. Accessing Limiting Reagent Homework Answer Key Free and Paid eBooks
 - Limiting Reagent Homework Answer Key Public Domain eBooks
 - Limiting Reagent Homework Answer Key eBook Subscription Services
 - Limiting Reagent Homework Answer Key Budget-Friendly Options
- 6. Navigating Limiting Reagent Homework Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reagent Homework Answer Key Compatibility with Devices
 - Limiting Reagent Homework Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reagent Homework Answer Key
 - Highlighting and Note-Taking Limiting Reagent Homework Answer Key
 - Interactive Elements Limiting Reagent Homework Answer Key
- 8. Staying Engaged with Limiting Reagent Homework Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reagent Homework Answer Key
- 9. Balancing eBooks and Physical Books Limiting Reagent Homework Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagent Homework Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reagent Homework Answer Key
 - Setting Reading Goals Limiting Reagent Homework Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reagent Homework Answer Key
 - Fact-Checking eBook Content of Limiting Reagent Homework Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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